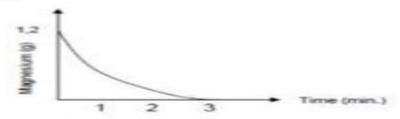
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records the change in the mass of magnesium as the reaction proceeds at regular intervals. The balanced chemical equation for the reaction is:

$$Mg(s) + 2HCl(aq) \rightarrow MgCl_2(aq) + H_2(g)$$

The change in the mass of magnesium during the reaction is shown on the graph below.



- 6.2.1 Identify the limiting agent in this reaction. Give a reason for the answer.
- 6.2.2 Calculate the number of moles of unreacted hydrochloric acid in the test tube after 3 minutes.

Hide Answers

QUESTION/VRAAG 6

- 6.1.1 2H₂O₂ (aq) ✓ → 2H₂O (ℓ) + O₂ (g) ✓
- 6.1.2 The catalyst lowers the activation energy of the reaction
 Accept: catalyst speeds up the reaction
 In Katalisator vertiage die aktiveringsenergie van die reaksie
 - 'n Katalisator verlaag die aktiveringsenergie van die reaksie 🗸 🗸 Aanvaar: katalisator laat die reaksie vinniger plaasvind
- 6.1.3 OPTION 1/OPSIE 1

$$0,049 = \frac{m}{34} \checkmark$$

OPTION 2/OPSIE 2
From the balanced equation:
Vanaf gebalanseerde
vergelyking:

68g $H_2O_2 \rightarrow 24,45 \text{ dm}^3 O_2 \checkmark \checkmark$ $X \text{ g } H_2O_2 \rightarrow 600 \times 10^{-3} \text{ dm}^3 \checkmark$ $X = 68 \times 0.6 \checkmark$ (2)

(7) [19]

(2)

$$X = \frac{68 \times 0.6}{24.45}$$

 $X = 1.67 g$

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