

QUESTION 3

3. 1.Learners made a mini volcano in science laboratory by adding sodium bicarbonate to ethanoic acid. They added 100 ml of a 0, 2 mol.dm⁻³ ethanoic acid solution to 10 g of NaHCO₃ to start the reaction of the volcano.

The balanced equation for this reaction is:



- 3.1.1 Define the term limiting reagent. (2)
- 3.1.2 Determine the limiting reagent in this reaction. (6)
- 3.1.3 Calculate the mass of the other substance in excess. (3)
- 3.1.4 Calculate the volume of CO₂ produced at STP. (4)

3.2 A solution of potassium hydroxide (KOH) is prepared by dissolving 3,36 g crystals of KOH in 250 cm³ of water.

Calculate the concentration of the potassium hydroxide solution. (4)

3.3 25 cm³ of a potassium hydroxide solution of concentration 0,25 mol.dm⁻³ completely neutralises a dilute solution of sulphuric acid (H₂SO₄) in a flask.

The unbalanced equation below represents the reaction that takes place:



- 3.3.1 Balance the above equation. (2)
- 3.3.2 Calculate the mass of sulphuric acid in the flask. (5)

[25]

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Thomas Griffiths



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level is 0.05, and the test is to compare the difference between two means ($\mu_1 - \mu_2$) against 0 ($H_0 \dots$