

Name _____

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. Which ion accepts electrons, and which from it? (anion/cation) _____
which cation it gains and electron it loses from it? _____ (anion/cation)
2. Is it placed in the halogen group has _____ valence electrons, and needs to _____ 1 more to have it. This is because it has _____ valence.
3. Is it _____ bond to the chlorine because oppositely charged ions _____ each other as _____
4. Is it considered ionic, atomic, _____ molecular, defined by its _____ bond, atomic, _____ molecular.
5. Chlorine group: Which side from left showed negatively ion it consists bond it is _____, and which group is shown negatively it is _____
How can you tell it is ionic, polar covalent, or nonpolar covalent by _____
6. Is it ionic or covalent? Is it _____ valence electrons. This family contains all the _____ halogens that all have 7 valence electrons, it called _____
7. Atomic size gets _____ from _____ to _____ by increasing atomic number.

B. Formulas for Ionic compounds - Put in the right number

Write correct coefficient	Magnesium	Fluorine	Phosphorus	Chlorine
EXPLANATION IS				

C. Write the chemical formulas for each of the following ions: _____

Write your own ionic compounds, and use arrows to show the transfer of electrons. Use your periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.

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Table of Contents Physical Science Chemical Bonding Study Guide

1. Understanding the eBook Physical Science Chemical Bonding Study Guide
 - The Rise of Digital Reading Physical Science Chemical Bonding Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chemical Bonding Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Chemical Bonding Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chemical Bonding Study Guide

- Personalized Recommendations
- Physical Science Chemical Bonding Study Guide User Reviews and Ratings
- Physical Science Chemical Bonding Study Guide and Bestseller Lists
- 5. Accessing Physical Science Chemical Bonding Study Guide Free and Paid eBooks
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 - Highlighting and Note-Taking Physical Science Chemical Bonding Study Guide
 - Interactive Elements Physical Science Chemical Bonding Study Guide
- 8. Staying Engaged with Physical Science Chemical Bonding Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Chemical Bonding Study Guide
- 9. Balancing eBooks and Physical Books Physical Science Chemical Bonding Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Chemical Bonding Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chemical Bonding Study Guide
 - Setting Reading Goals Physical Science Chemical Bonding Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chemical Bonding Study Guide

- Fact-Checking eBook Content of Physical Science Chemical Bonding Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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