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4.14.

$$\left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.000000$$

$$\text{mass of H}_2\text{O} = 1.000000 \text{ g} \times 1000 \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1000 \text{ g}$$

and that would involve 1000 g of hydrogen.

- 4.15. (a) The nuclei involved here have the same mass number, 12.
 (b) The nuclei involved have the same mass number, 12.
 (c) The nuclei involved have the same mass number, 12.

4.16. (a) The mass of the nucleus is the sum of the masses of the protons and neutrons.
 (b) The mass of the nucleus is the sum of the masses of the protons and neutrons.

$$\text{mass of } ^{12}\text{C} = 12.000000 \text{ u} \times \left(\frac{1.660539 \times 10^{-27} \text{ kg}}{1.660539 \times 10^{-27} \text{ kg}} \right) = 1.992647 \times 10^{-26} \text{ kg}$$

4.17. (a) The mass of the nucleus is the sum of the masses of the protons and neutrons.

$$\text{mass of } ^{12}\text{C} = 12.000000 \text{ u} \times \left(\frac{1.660539 \times 10^{-27} \text{ kg}}{1.660539 \times 10^{-27} \text{ kg}} \right) = 1.992647 \times 10^{-26} \text{ kg}$$

$$\text{mass of } ^{12}\text{C} = 12.000000 \text{ u} \times \left(\frac{1.660539 \times 10^{-27} \text{ kg}}{1.660539 \times 10^{-27} \text{ kg}} \right) = 1.992647 \times 10^{-26} \text{ kg}$$

4.18. (a) The mass of the nucleus is the sum of the masses of the protons and neutrons.

$$\text{mass of } ^{12}\text{C} = 12.000000 \text{ u} \times \left(\frac{1.660539 \times 10^{-27} \text{ kg}}{1.660539 \times 10^{-27} \text{ kg}} \right) = 1.992647 \times 10^{-26} \text{ kg}$$

Physical Chemistry Ball Solution Manual

Peipei Pang



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