Chapter 1

Some Basic Concepts of Chemistry

Question 1.1:

Calculate the molar mass of the following:

(i) H₂O (ii) CO₂ (iii) CO₄

Answer 1.1:

(i) H,O:

The molecular mass of water, H₂O

- = (2 × Atomic mass of hydrogen) + (1 × Atomic mass of oxygen)
- = [2(1.0084) + 1(16.00 u)]
- = 2.016 u + 16.00 u = 18.016 = 18.02 u
- (iii) CO₂:

The molecular mass of carbon dioxide, CO,

- = (1 × Atomic mass of carbon) + (2 × Atomic mass of oxygen)
- = [1(12.011 u) + 2(16.00 u)]
- = 12.011 u + 32.00 u = 44.01 u
- (iiii) CH,:

The molecular mass of methane, CH.

- = (1 × Atomic mass of carbon) + (4 × Atomic mass of hydrogen)
- = [1(12.011 u) + 4(1.008 u)]
- = 12.011 u + 4.032 u = 16.043 u

Question 1.2:

Calculate the mass per cent of different elements present in sodium sulphate (Na₂SO₄).

Answer 1.2:

The molecular formula of sodium sulphate is (Na₂SO₄).

Molar mass of $Na_2SO_4 = [(2 \times 23.0) + (32.066) + 4 (16.00)] = 142.066 g$

Mass percent of an element = $\frac{Mass \ of \ that \ element \ in \ the \ compound}{Molar \ Mass \ of \ the \ compond} \times 100$

... Mass percent of sodium: $\frac{46.0g}{142.066g} \times 100 = 32.379 = 32.4\%$

Mass percent of sulphur: $\frac{32.066g}{142.066g} \times 100 = 22.57 = 22.6\%$

Mass percent of oxygen: $\frac{46.0g}{142.066g} \times 100 = 42.049 = 45.05\%$

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