

- (i) Express this in percent by mass**
(ii) Determine the molality of chloroform in the water sample.

Answer:

(i) 15 ppm means 15 parts in million (10^6) parts

$$\therefore \% \text{ by mass} = \frac{15}{10^6} \times 100 = 15 \times 10^{-4} = 1.5 \times 10^{-3} \%$$

(ii) Molar mass of chloroform (CHCl_3) = $12 + 1 + 3 \times 35.5 = 119.5 \text{ g mol}^{-1}$

100 g of the sample contain chloroform = $1.5 \times 10^{-3} \text{ g}$

\therefore 1000 g (1 kg) of the sample will contain chloroform = $1.5 \times 10^{-2} \text{ g}$

$$= \frac{1.5 \times 10^{-2}}{119.5} = 1.26 \times 10^{-4} \text{ mole}$$

\therefore Molality = $1.266 \times 10^{-4} \text{ m}$.

Question 18. Express the following in scientific notation:

(i) 0.0048 (v) 6.0012 (ii) 234,000 (iii) 8008 (iv) 500.0

Answer:

(i) 4.8×10^{-3}

(ii) 2.34×10^5

(iii) 8.008×10^3

(iv) 5.000×10^2

(v) 6.0012×10^0

Question 19. How many significant figures are present in the following?

(i) 0.0025 (ii) 208 (iii) 5005 (iv) 126,000

(v) 500.0 (vi) 2.0034

Answer: (i) 2 (ii) 3 (iii) 4 (iv) 3 (v) 4 (vi) 5.

Question 20. Round up the following upto three significant figures:

(i) 34.216 (ii) 10.4107 (iii) 0.04597 (iv) 2808

Answer: (i) 34.2 (ii) 10.4 (iii) 0.0460 (iv) 2810

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