

## Chapter 6: Chemical Bonding Guided Notes

### I. Why do Atoms Combine?

#### A. Chemical Formula

1. The chemical formula shows:

- The \_\_\_\_\_ in the compound.
- The ratio of their \_\_\_\_\_.



#### B. Chemical Bond

1. Strong attractive force between \_\_\_\_\_ or ions in a molecule or compound.

2. Formed by:

- electrons, ( $e^-$ ) by losing or gaining electrons.
- electrons ( $e^-$ )

#### C. Stability

1. Octet Rule

- Most atoms form bonds in order to have \_\_\_\_\_ valence electrons ( $e^-$ ).
- outer energy level.
- Want to be like the Noble Gases (group 18).



2. Stability is the driving force behind bond formation. All atoms bond to become stable.

3. Transferring  $e^-$ :



4. Sharing  $e^-$ :



### II. Kinds of Chemical Bonds

#### A. Ionic Bond

1. Attraction between 2 \_\_\_\_\_ charged ions.



2. Ions are \_\_\_\_\_ atoms.

3. A \_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom loses an electron.

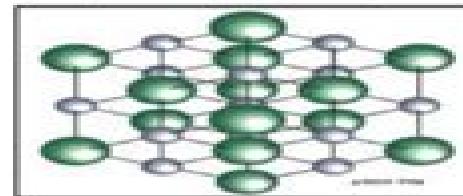
Formed by \_\_\_\_\_.

4. A \_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom gains an electron.

Formed by \_\_\_\_\_.

5. Formed by \_\_\_\_\_ electrons ( $e^-$ ) from a \_\_\_\_\_ to a \_\_\_\_\_.

6. Ions form a 3-D crystal lattice.



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## **Physical Science Chapter 6 Chemical Bonding Wordwise Answers**

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