

**QUESTION 10 (Start on a new page.)**

The balanced chemical equation for the EXTRACTION of gold from its ore is as follows:



- 10.1 State ONE disadvantage of using cyanide ( $\text{CN}^-$ ) in the extraction of gold. (1)
- 10.2 Will the final solution of the extraction process be ACIDIC or BASIC (ALKALINE)? Give a reason for the answer. (2)
- 10.3 Determine the oxidation number of gold in  $\text{NaAu(CN)}_2$ . (1)
- 10.4 Write down the FORMULA of the reducing agent in the reaction above. (1)

Zinc powder is now used to PRECIPITATE the gold.

The balanced equation for the reaction is:



- 10.5 Does zinc undergo OXIDATION or REDUCTION during the precipitation reaction? (1)
- 10.6 Write down a half-reaction to support the answer to QUESTION 10.5. (2)
- 10.7 Calculate the percentage of gold in  $\text{NaAu(CN)}_2$ . (2)
- [10]**

**TOTAL: 150**

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## **Table of Contents Physical Science P2 Caps Grade1dbe November 2014**

1. Understanding the eBook Physical Science P2 Caps Grade1dbe November 2014
  - The Rise of Digital Reading Physical Science P2 Caps Grade1dbe November 2014
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science P2 Caps Grade1dbe November 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science P2 Caps Grade1dbe November 2014
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science P2 Caps Grade1dbe November 2014
  - Personalized Recommendations
  - Physical Science P2 Caps Grade1dbe November 2014 User Reviews and Ratings
  - Physical Science P2 Caps Grade1dbe November 2014 and Bestseller Lists

5. Accessing Physical Science P2 Caps Grade1dbe November 2014 Free and Paid eBooks
  - Physical Science P2 Caps Grade1dbe November 2014 Public Domain eBooks
  - Physical Science P2 Caps Grade1dbe November 2014 eBook Subscription Services
  - Physical Science P2 Caps Grade1dbe November 2014 Budget-Friendly Options
6. Navigating Physical Science P2 Caps Grade1dbe November 2014 eBook Formats
  - ePub, PDF, MOBI, and More
  - Physical Science P2 Caps Grade1dbe November 2014 Compatibility with Devices
  - Physical Science P2 Caps Grade1dbe November 2014 Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science P2 Caps Grade1dbe November 2014
  - Highlighting and Note-Taking Physical Science P2 Caps Grade1dbe November 2014
  - Interactive Elements Physical Science P2 Caps Grade1dbe November 2014
8. Staying Engaged with Physical Science P2 Caps Grade1dbe November 2014
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science P2 Caps Grade1dbe November 2014
9. Balancing eBooks and Physical Books Physical Science P2 Caps Grade1dbe November 2014
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science P2 Caps Grade1dbe November 2014
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physical Science P2 Caps Grade1dbe November 2014
  - Setting Reading Goals Physical Science P2 Caps Grade1dbe November 2014
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science P2 Caps Grade1dbe November 2014
  - Fact-Checking eBook Content of Physical Science P2 Caps Grade1dbe November 2014
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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