

## Blackline Master 5.8

### Ionic Compounds: Names and Formulas Worksheet

1. Write the formulas for the following compounds.

|                                     |                                    |
|-------------------------------------|------------------------------------|
| (a) magnesium oxide                 | <u>MgO</u>                         |
| (b) sodium fluoride                 | <u>NaF</u>                         |
| (c) aluminum nitride                | <u>AlN</u>                         |
| (d) potassium sulfide               | <u>K<sub>2</sub>S</u>              |
| (e) lithium iodide                  | <u>LiI</u>                         |
| (f) calcium bromide                 | <u>CaBr<sub>2</sub></u>            |
| (g) beryllium oxide                 | <u>BeO</u>                         |
| (h) nickel <sup>(II)</sup> chloride | <u>NiCl<sub>2</sub></u>            |
| (i) magnesium nitride               | <u>Mg<sub>3</sub>N<sub>2</sub></u> |
| (j) aluminum sulfide                | <u>Al<sub>2</sub>S<sub>3</sub></u> |

|                        |                                    |
|------------------------|------------------------------------|
| (k) copper(I) bromide  | <u>CuBr</u>                        |
| (l) tin(II) iodide     | <u>SnI<sub>2</sub></u>             |
| (m) iron(III) chloride | <u>FeCl<sub>3</sub></u>            |
| (n) calcium phosphide  | <u>Ca<sub>3</sub>P<sub>2</sub></u> |
| (o) lead(II) oxide     | <u>PbO</u>                         |
| (p) lead(IV) fluoride  | <u>PbF<sub>4</sub></u>             |
| (q) tin(IV) bromide    | <u>SnBr<sub>4</sub></u>            |
| (r) copper(II) sulfide | <u>CuS</u>                         |
| (s) iron(II) oxide     | <u>FeO</u>                         |
| (t) calcium nitride    | <u>Ca<sub>3</sub>N<sub>2</sub></u> |

2. Write the names for the following compounds.

|                                    |                            |
|------------------------------------|----------------------------|
| (a) Li <sub>2</sub> O              | <u>lithium oxide</u>       |
| (b) AlCl <sub>3</sub>              | <u>aluminum chloride</u>   |
| (c) MgS                            | <u>magnesium sulfide</u>   |
| (d) CaO                            | <u>calcium oxide</u>       |
| (e) KBr                            | <u>potassium bromide</u>   |
| (f) BeF                            | <u>beryllium fluoride</u>  |
| (g) Na <sub>3</sub> N              | <u>sodium nitride</u>      |
| (h) Al <sub>2</sub> O <sub>3</sub> | <u>aluminum oxide</u>      |
| (i) CuCl <sub>2</sub>              | <u>copper(II) chloride</u> |
| (j) FeBr <sub>3</sub>              | <u>iron(III) bromide</u>   |

|                                    |                             |
|------------------------------------|-----------------------------|
| (k) PbS                            | <u>lead (II) sulfide</u>    |
| (l) SnO <sub>2</sub>               | <u>tin (IV) oxide</u>       |
| (m) Na <sub>2</sub> S              | <u>sodium sulfide</u>       |
| (n) Mg <sub>3</sub> P <sub>2</sub> | <u>magnesium phosphide</u>  |
| (o) NiO                            | <u>nickel (II) oxide</u>    |
| (p) CuI                            | <u>copper (I) iodide</u>    |
| (q) PbCl <sub>2</sub>              | <u>lead (II) chloride</u>   |
| (r) FeP                            | <u>iron (III) phosphide</u> |
| (s) CaF <sub>2</sub>               | <u>calcium fluoride</u>     |
| (t) K <sub>3</sub> P               | <u>potassium phosphide</u>  |

# Physical Science Names And Formulas Worksheet 1

**JA Banks**



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