# **Answers to questions**

## Chapter 1

- 1.1 Petroleum and natural gas are formed from the bodies of dead marine creatures subjected to heat and pressure over very long geological periods of time.
- 1.2 It provides the energy needed for photosynthesis, which removes carbon dioxide from the atmosphere.
- 1.3 They are regarded as non-renewable because they were formed over very long periods of time and are being used up at a rate far faster than they can be formed.
- 1.4 a carbon dioxide + water → glucose + oxygen (chemical equation) 6CO<sub>2</sub> + 6H<sub>2</sub>O → C<sub>s</sub>H<sub>12</sub>O<sub>s</sub> + 6O<sub>2</sub>
  - b carbon + oxygen → carbon dioxide (chemical equation) C + O<sub>1</sub> → CO<sub>2</sub>
  - c glucose + oxygen  $\rightarrow$  carbon dioxide + water

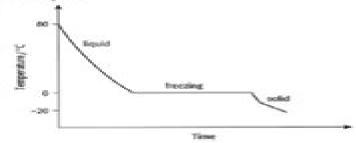
(chemical equation)  $C_aH_{12}O_a + 6O_2$  $\rightarrow 6CO_2 + 6H_2O_3$ 

- 1.5 sulfur dioxide and nitrogen dioxide
- 1.6 the burning of fossil fuels (mainly coal for sulfur dioxide)
- 1.7 damage to limestone buildings, death of trees, acidification of lakes leading to death of fish
- 1.8 a combination of nitrogen oxides and low-level ozone that causes breathing problems, especially for people with asthma
- 1.9 It combines with the haemoglobin in red blood cells, stopping them from carrying oxygen.
- 1.10 because it does not react with the filament, which would burn in air when it became hot
- 1.11 from rotting vegetable matter and from the intestines of animals such as coves
- 1.12 Heat which would normally escape into space is reflected back to the Earth's surface by gases such as carbon dioxide and methane in the atmosphere.
- 1.13 It changes nitrogen oxides and carbon monoxide to nitrogen and carbon dioxide.
- 1.14 because they have different boiling points
- 1.15 because solid matter is easiest to remove and would interfere with subsequent processes

- 1.16 to kill bacteria present in the water
- 1.17 because the energy needed to boil the water is costly
- 1.18 The sewage reacts with oxygen in the water, leaving less for water creatures to breathe.
- 1.19 They can cause too many algae to grow and they are not removed by water treatment.
- It contains a high concentration of a particular metal compound.
- Lime is calcium oxide which is reacted with water to make calcium hydroxide (slaked lime).
- 1.22 plastics, drugs, paints, detergents, etc.
- 1.23 It reflects heat back to Earth when present in the atmosphere. Heat is kept in the atmospheric layer.
- 1.24 The electrolysis of water. It is an expensive process.
- 1.25 It produces only water when burned.
- 1.26 hydrogen + oxygen → water 2H<sub>1</sub> + O<sub>2</sub> → 2H<sub>2</sub>O

## Chapter 2

- 2.1 a freezing (solidification)
  - b boiling
  - c condensation
  - d sublimation
- 2.2 The impurity lowers the freezing point of the liquid.
- 2.3



2.4 A volatile liquid is one that evaporates easily; it has a low boiling point.

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**David Kirk** 

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