

## Chapter 1

### Some Basic Concepts of Chemistry

#### Answer 1.9:

The average atomic mass of chlorine

$$\begin{aligned}
 &= \left[ (\text{Fractional abundance of } ^{35}\text{Cl}) (\text{Molar mass of } ^{35}\text{Cl}) \right] \\
 &+ \left[ (\text{Fractional abundance of } ^{37}\text{Cl}) (\text{Molar mass of } ^{37}\text{Cl}) \right] \\
 &= \left[ \left\{ \left( \frac{75.77}{100} \right) (34.9689\text{u}) \right\} + \left\{ \left( \frac{24.23}{100} \right) (36.9659\text{u}) \right\} \right] \\
 &= 26.4959 + 8.9568 = 35.4527 \text{ u}
 \end{aligned}$$

The average atomic mass of chlorine = 35.4527 u

#### Question 1.10:

In three moles of ethane ( $\text{C}_2\text{H}_6$ ), calculate the following:

1. Number of moles of carbon atoms.
2. Number of moles of hydrogen atoms.
3. Number of molecules of ethane.

#### Answer 1.10:

1. 1 mole of  $\text{C}_2\text{H}_6$  contains 2 moles of carbon atoms.  
 $\therefore$  Number of moles of carbon atoms in 3 moles of  $\text{C}_2\text{H}_6 = 2 \times 3 = 6$
2. 1 mole of  $\text{C}_2\text{H}_6$  contains 6 moles of hydrogen atoms.  
 $\therefore$  Number of moles of carbon atoms in 3 moles of  $\text{C}_2\text{H}_6 = 3 \times 6 = 18$
3. 1 mole of  $\text{C}_2\text{H}_6$  contains  $6.023 \times 10^{23}$  molecules of ethane.  
 $\therefore$  Number of molecules in 3 moles of  $\text{C}_2\text{H}_6$   
 $= 3 \times 6.023 \times 10^{23} = 18.069 \times 10^{23}$

#### Question 1.11:

What is the concentration of sugar ( $\text{C}_{12}\text{H}_{22}\text{O}_{11}$ ) in  $\text{mol L}^{-1}$  if its 20g are dissolved in enough water to make a final volume up to 2L.

#### Answer 1.11:

Molarity (M) of a solution is given by,

$$\begin{aligned}
 &= \frac{\text{Number of moles of solute}}{\text{Volume of solution in Litres}} = \frac{\text{Mass of sugar/molar mass of sugar}}{2\text{L}} \\
 &= \frac{20\text{g}/[(12 \times 12) + (1 \times 22) + (11 \times 16)] \text{ g}}{2\text{L}} = \frac{20\text{g}/342 \text{ g}}{2\text{L}} = \frac{0.0585 \text{ mol}}{2\text{L}} = 0.02925 \text{ mol L}^{-1}
 \end{aligned}$$

$\therefore$  Molar concentration of sugar = 0.02925  $\text{mol L}^{-1}$

# Ncert 11 Chemistry Chapter Wise Notes

**Clifford Lane**



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### **Table of Contents Ncert 11 Chemistry Chapter Wise Notes**

1. Understanding the eBook Ncert 11 Chemistry Chapter Wise Notes
  - The Rise of Digital Reading Ncert 11 Chemistry Chapter Wise Notes
  - Advantages of eBooks Over Traditional Books
2. Identifying Ncert 11 Chemistry Chapter Wise Notes
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Ncert 11 Chemistry Chapter Wise Notes
  - User-Friendly Interface
4. Exploring eBook Recommendations from Ncert 11 Chemistry Chapter Wise Notes
  - Personalized Recommendations
  - Ncert 11 Chemistry Chapter Wise Notes User Reviews and Ratings
  - Ncert 11 Chemistry Chapter Wise Notes and Bestseller Lists
5. Accessing Ncert 11 Chemistry Chapter Wise Notes Free and Paid eBooks

- Ncert 11 Chemistry Chapter Wise Notes Public Domain eBooks
  - Ncert 11 Chemistry Chapter Wise Notes eBook Subscription Services
  - Ncert 11 Chemistry Chapter Wise Notes Budget-Friendly Options
6. Navigating Ncert 11 Chemistry Chapter Wise Notes eBook Formats
- ePub, PDF, MOBI, and More
  - Ncert 11 Chemistry Chapter Wise Notes Compatibility with Devices
  - Ncert 11 Chemistry Chapter Wise Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Ncert 11 Chemistry Chapter Wise Notes
  - Highlighting and Note-Taking Ncert 11 Chemistry Chapter Wise Notes
  - Interactive Elements Ncert 11 Chemistry Chapter Wise Notes
8. Staying Engaged with Ncert 11 Chemistry Chapter Wise Notes
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Ncert 11 Chemistry Chapter Wise Notes
9. Balancing eBooks and Physical Books Ncert 11 Chemistry Chapter Wise Notes
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Ncert 11 Chemistry Chapter Wise Notes
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Ncert 11 Chemistry Chapter Wise Notes
- Setting Reading Goals Ncert 11 Chemistry Chapter Wise Notes
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ncert 11 Chemistry Chapter Wise Notes
- Fact-Checking eBook Content of Ncert 11 Chemistry Chapter Wise Notes
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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