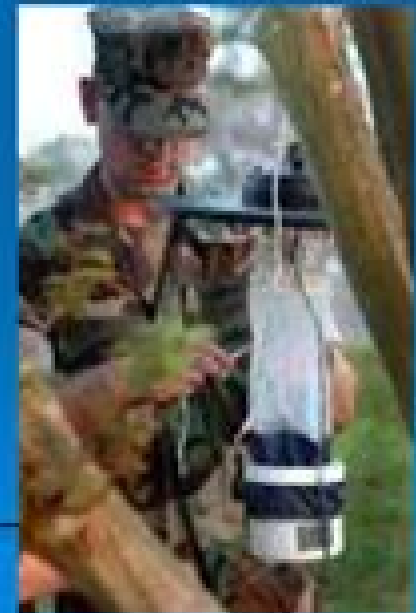


NAVMED P-5010 CHAPTER 9 PREVENTIVE MEDICINE FOR GROUND FORCES

Presented By:
HM1(SW) Marlene Rivera



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Chapter 9

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Tactical Combat Casualty Care TCCC was developed to emphasize the need for continued improvement in combat pre hospital care The Committee on Tactical Combat Casualty Care CoTCCC was established in 2001 and is part of the Defense Health Board CoTCCC is a standing multi service committee charged with monitoring medical developments in regards to practice technology pharmacology and doctrine New concepts in hemorrhage control airway management fluid resuscitation analgesia antibiotics and other lifesaving techniques are important steps in providing the best possible care for our Marines and Sailors in combat The TCCC guidelines are published every 4 years in the Prehospital Trauma Life Support manual It has been recognized that TCCC guidelines and curriculum will need to change more often than the 4 year cycle of the PHTLS textbook publication The National Association of Emergency Medical Technicians NAEMT will include the updated TCCC guidelines and curriculum on its website as they are approved as a way to help get this new information out to the combat medical personnel in the military that need it PRINCIPLES OF TACTICAL COMBAT CASUALTY CARE TCCC The principles of Tactical Combat Casualty Care are fundamentally different from those of traditional civilian trauma care where most medical providers and medics train These differences are based on both the unique patterns and types of wounds that are suffered in combat and the tactical conditions medical personnel face in combat Unique combat wounds and tactical conditions make it difficult to determine which intervention to perform at what time Besides addressing a casualty's medical condition responding medical personnel must also address the tactical problems faced while providing care in combat A medically correct intervention at the wrong time may lead to further casualties Put another way good medicine may be a bad tactical decision which can get the rescuer and the casualty killed To successfully navigate these issues medical providers must have skills and training oriented to combat trauma care as opposed to civilian trauma care The specifics of casualty care in the tactical setting will depend on the tactical situation the injuries sustained by the casualty the knowledge and skills of the first responder and the medical equipment at hand In contrast to a hospital Emergency Department setting where the patient IS the mission on the battlefield care of casualties sustained is only PART of the mission TCCC recognizes this fact and structures its guidelines to accomplish three primary goals 1 Treat the casualty 2 Prevent additional casualties 3 Complete

the mission In thinking about the management of combat casualties it is helpful to divide care into three distinct phases each with its own characteristics and limitations **Vector Control** United States. Navy Department. Bureau of Medicine and Surgery,1958 Mess Management Specialist 1 & C. United States. Naval Education and Training Command,1978

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web atoms of the same element i.e. same atomic number Z that have different numbers of neutrons are called isotopes for example 99 of the carbon atoms on earth have 6 neutrons and 6 protons in their nuclei about 1 of the carbon atoms have 7

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web aug 19 2022 atoms with the same number of protons but different numbers of neutrons are called isotopes they share almost the same chemical properties but differ in mass and therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation

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web aug 9 2023 how do different isotopes of an element differ two isotopes differ in the number of neutrons in the nucleus for example carbon 14 has 8 neutrons differing from carbon 12 which has 6

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web atoms of the same element that differ in their numbers of neutrons are called isotopes many isotopes occur naturally

usually one or two isotopes of an element are the most stable and common different isotopes of an element generally have the same physical and chemical properties because they have the same numbers of protons and electrons

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web jan 5 2016 explanation isotopes are atoms of the same element same atomic number ie same number of protons in the nucleus but different atomic masses different mass numbers due to different numbers of neutrons in the nucleus for example we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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web may 16 2018 explanation an atom is characterized as a particular element by the number of positively charged massive nuclear particles i e z $z=1$ we get hydrogen $z=2$ we get helium $z=3$ lithium $z=4$ we get silver the nucleus can also contain neutrons massive nuclear particles of zero charge

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web atoms of the same element with different numbers of neutrons are called isotopes close isotope atoms of an element with the same number of protons and electrons but different numbers of neutrons

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web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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web jul 18 2018 answer link due to the number of neutrons isotopes differ with respect to the number of neutrons their nuclei contain where neutrons are massive particles of zero charge a nucleus may contain variable numbers of neutrons and this gives rise to the phenomenon of isotopes

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web isotopes are distinct nuclear species or nuclides as technical term of the same chemical element they have the same atomic number number of protons in their nuclei and position in the periodic table and hence belong to the same chemical element but differ in nucleon numbers mass numbers due to different numbers of neutrons in

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web verified by toppr correct option is d atoms having same atomic number but different atomic mass are known as isotopes the difference in the mass is due to the difference in their number of neutrons with the same number of electrons and protons

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope s mass number

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web nov 27 2016 explanation basically isotopes are nuclei that have the same number of protons and are thus the same element but have unequal numbers of neutrons answer link the nucleus differs as a result of the number of neutrons in the atom the actual structure doesn t differ all that much but its mass does differ thus its density is a little

solved isotopes differ from each other in what ways a they - Mar 04 2022

web chemistry questions and answers isotopes differ from each other in what ways a they have different numbers of electrons outside the nucleus b they have different atomic number o c they have different numbers of neutrons in the nucleus d more than one response is correct e they have different numbers of protons in the nucleus