

## Unit 7: Balancing Chemical Reactions • Worksheet 2

Balance the following equations by inserting the proper coefficients.

1.  $\underline{2}\text{SO}_2 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{SO}_3$
  2.  $\underline{2}\text{CH}_4 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO} + \underline{4}\text{H}_2\text{O}$
  3.  $\underline{2}\text{P} + \underline{3}\text{Cl}_2 \rightarrow \underline{2}\text{PCl}_3$
  4.  $\underline{2}\text{CO} + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CO}_2$
  5.  $\underline{2}\text{CH}_4 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CH}_3\text{OH}$
  6.  $\underline{2}\text{Li} + \underline{1}\text{Br}_2 \rightarrow \underline{2}\text{LiBr}$
  7.  $\underline{2}\text{Al}_2\text{O}_3 \rightarrow \underline{4}\text{Al} + \underline{3}\text{O}_2$
  8.  $\underline{2}\text{Na} + \underline{2}\text{H}_2\text{O} \rightarrow \underline{2}\text{NaOH} + \underline{1}\text{H}_2$
  9.  $\underline{6}\text{CO}_2 + \underline{6}\text{H}_2\text{O} \rightarrow \underline{1}\text{C}_6\text{H}_{12}\text{O}_6 + \underline{6}\text{O}_2$
  10.  $\underline{1}\text{H}_2\text{SO}_4 + \underline{2}\text{NaCl} \rightarrow \underline{2}\text{HCl} + \underline{1}\text{Na}_2\text{SO}_4$
- Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:
11.  $\underline{3}\text{H}_2 + \underline{1}\text{SO}_2 \rightarrow \underline{1}\text{H}_2\text{S} + \underline{2}\text{H}_2\text{O}$
  12.  $\underline{2}\text{CaCO}_3 + \underline{2}\text{SO}_2 + \underline{1}\text{O}_2 \rightarrow \underline{2}\text{CaSO}_4 + \underline{2}\text{CO}_2$
- 
13.  $\underline{2}\text{AgNO}_3 + \underline{1}\text{CaCl}_2 \rightarrow \underline{2}\text{AgCl} + \underline{1}\text{Ca}(\text{NO}_3)_2$
  14.  $\underline{2}\text{HCl} + \underline{1}\text{Ba}(\text{OH})_2 \rightarrow \underline{1}\text{BaCl}_2 + \underline{2}\text{H}_2\text{O}$
  15.  $\underline{1}\text{H}_3\text{PO}_4 + \underline{3}\text{NaOH} \rightarrow \underline{1}\text{Na}_3\text{PO}_4 + \underline{3}\text{H}_2\text{O}$
  16.  $\underline{1}\text{Pb}(\text{NO}_3)_2 + \underline{2}\text{KI} \rightarrow \underline{1}\text{PbI}_2 + \underline{2}\text{KNO}_3$
  17.  $\underline{3}\text{CuO} + \underline{2}\text{NH}_3 \rightarrow \underline{1}\text{N}_2 + \underline{3}\text{Cu} + \underline{3}\text{H}_2\text{O}$
  18.  $\underline{1}\text{C}_2\text{H}_5\text{OH} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CO}_2 + \underline{3}\text{H}_2\text{O}$
  19.  $\underline{2}\text{C}_2\text{H}_6 + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{CH}_3\text{COOH} + \underline{2}\text{H}_2\text{O}$
  20.  $\underline{3}\text{NO}_2 + \underline{1}\text{H}_2\text{O} \rightarrow \underline{2}\text{HNO}_3 + \underline{1}\text{NO}$

# Rearranging Atoms Worksheet

**Debbie Lawrence, Richard Lawrence**

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

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