Appendix D ANSWERS TO SOME PROBLEMS

1-1. (a) At standard temperature and pressure, a mole of an ideal gas contains 6.022×10^{28} molecules (Avogadro's number) and occupies 22.4 liters. Hence, the number per m³ is $6.022 \times 10^{28}/2.24 \times 10^{-8} = 2.66 \times 10^{28} \, \mathrm{m}^{-8}$.

(b) Since PV = NRT, n = N/V = P/RT. Hence $n_1/n_0 = P_1T_0/P_0T_1$. Taking n_0 to be the density in part (a) and n_1 to be that in part (b), we have

$$n_1 = (2.69 \times 10^{20}) \frac{10^{-9}}{760} \frac{273}{(273 + 20)} = 3.30 \times 10^{19} \,\mathrm{m}^{-8}$$

Note that a diatomic gas such as H_0 will have twice as many atoms per torr as, say, He.

1-2. Consider the integral

$$I^{2} = \int_{-\infty}^{\infty} e^{-x^{2}} dx \int_{-\infty}^{\infty} e^{-y^{2}} dy = \int_{-\infty}^{\infty} e^{-(x^{2} + y^{2})} dx dy$$

in a two-dimensional space. Transforming to cylindrical coordinates, we have

$$I^{\alpha} = \int \int e^{-r^{\alpha}} r \, dr \, d\phi = 2\pi \int_{0}^{\infty} e^{-r^{\alpha}} r \, dr$$

= $= \int e^{-r^{\alpha}} d(r^{\alpha}) = -\pi e^{-r^{\alpha}} |_{0}^{\infty} = \pi$

Hence,

1

$$I = \int_{-\infty}^{\infty} e^{-x^2} dx = \sqrt{\pi}$$

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Table of Contents Plasma Chen Solution Manual

- 1. Understanding the eBook Plasma Chen Solution Manual
 - The Rise of Digital Reading Plasma Chen Solution Manual
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Plasma Chen Solution Manual
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Plasma Chen Solution Manual
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Plasma Chen Solution Manual
 - Personalized Recommendations

- Plasma Chen Solution Manual User Reviews and Ratings
- Plasma Chen Solution Manual and Bestseller Lists
- 5. Accessing Plasma Chen Solution Manual Free and Paid eBooks
 - Plasma Chen Solution Manual Public Domain eBooks
 - Plasma Chen Solution Manual eBook Subscription Services
 - Plasma Chen Solution Manual Budget-Friendly Options
- 6. Navigating Plasma Chen Solution Manual eBook Formats
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 - Interactive Elements Plasma Chen Solution Manual
- 8. Staying Engaged with Plasma Chen Solution Manual
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - o Following Authors and Publishers Plasma Chen Solution Manual
- 9. Balancing eBooks and Physical Books Plasma Chen Solution Manual
 - \circ Benefits of a Digital Library
 - Creating a Diverse Reading Collection Plasma Chen Solution Manual
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Plasma Chen Solution Manual
 - Setting Reading Goals Plasma Chen Solution Manual
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Plasma Chen Solution Manual
 - Fact-Checking eBook Content of Plasma Chen Solution Manual

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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