


		<p>gerangskik ☑ is om die hoeke tussen hulle te maksimeer. ☑ Strukturele elektronpare is bindingspare plus alleenpare.</p> <p>OF</p> <p>Valensdop-elektronpaarafstotingsmodel is 'n model om die vorms van molekules te voorspel ☑ waarin die elektronpaar vanaf die buitenste dop ☑ van 'n verwysingsatoom rondom hierdie atoom gerangskik is om die afstoting tussen hulle te minimaliseer. ☑</p> <p>3 merks/punte</p>	
1.2.4		<p>The molecule is a tetrahedron ☑ Die molekule is 'n tetraëder ☑</p> <p>1 merk/punt</p>	LO2
1.2.5		 <p>Name trigonal pyramid ☑ Noem trigonale piramide</p> <p>3 merks/punte</p>	LO4
1.3	1.3.1	<p>The electron pair is equally shared between the two Cl atoms ☑ and the resulting Cl₂ molecule has no degree of polarity. ☑ Die elektronpaar word ewe veel gedeel tussen die twee Cl-atome ☑ en die resulterende Cl₂-molekule het geen mate van polariteit nie ☑.</p> <p>2 merks/punte</p>	LO2
	1.3.2	<p>The tendency ☑ of an atom to pull the bonded pair ☑ of electrons towards itself. ☑ Die neiging ☑ van 'n atoom om die gebinde elektronpare ☑ na homself te trek ☑.</p> <p>3 merks/punte</p>	LO2
	1.3.3	<p>Electronegativity of Na = 6 Electronegativity of Cl = 3 Difference is 3 ☑</p> <p>With a difference of greater than 2.1 ☑ electron transfer will take place and the bond ☑ is ionic. ☑ (value of 2.1 does not need to be given for this mark) Elektronegatiwiteit van Na = 6 Elektronegatiwiteit van Cl = 3 Verskil is 3 ☑</p> <p>Met 'n verskil van meer as 2,1 ☑ sal elektroneoordrag plaasvind en die binding is ionies. ☑ (waarde van 2.1 hoef nie vir hierdie punt gegee te word nie)</p> <p>4 merks/punte</p>	LO2

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Disha Experts



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Table of Contents Physical Science 2014 Control Test

1. Understanding the eBook Physical Science 2014 Control Test
 - The Rise of Digital Reading Physical Science 2014 Control Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science 2014 Control Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science 2014 Control Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science 2014 Control Test
 - Personalized Recommendations

- Physical Science 2014 Control Test User Reviews and Ratings
- Physical Science 2014 Control Test and Bestseller Lists
- 5. Accessing Physical Science 2014 Control Test Free and Paid eBooks
 - Physical Science 2014 Control Test Public Domain eBooks
 - Physical Science 2014 Control Test eBook Subscription Services
 - Physical Science 2014 Control Test Budget-Friendly Options
- 6. Navigating Physical Science 2014 Control Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Physical Science 2014 Control Test Compatibility with Devices
 - Physical Science 2014 Control Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science 2014 Control Test
 - Highlighting and Note-Taking Physical Science 2014 Control Test
 - Interactive Elements Physical Science 2014 Control Test
- 8. Staying Engaged with Physical Science 2014 Control Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science 2014 Control Test
- 9. Balancing eBooks and Physical Books Physical Science 2014 Control Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science 2014 Control Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science 2014 Control Test
 - Setting Reading Goals Physical Science 2014 Control Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science 2014 Control Test
 - Fact-Checking eBook Content of Physical Science 2014 Control Test

-
- Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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