

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- math

Kinetic Theory and Boyle's Law

- Pressure of a gas is caused by the collisions of the gas molecules impinging on the walls.
- If the gas is compressible to small the volume it has less surface area and the pressure is greater with less surface.
- Force is mass times acceleration per area on the walls of the container.
- Pressure of the gas will increase.

Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. After the volume is reduced, it maintains a pressure of 1.25 atm above the pressure inside the container. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 1.01 \text{ atm} \quad P_2 = 1.01 + 1.25 \text{ atm} = 2.26 \text{ atm}$$

$$V_2 = \frac{P_1 V_1}{P_2} = \frac{800 \text{ ml} \times 1.01 \text{ atm}}{2.26 \text{ atm}} = 358 \text{ ml}$$

Ex. 2. A gas is pressure of 1.01 atm and has a volume of 5.5 L. If the pressure is increased until the volume is 3.0 L, what is the new pressure, assuming the temperature remains constant?

$$P_1 V_1 = P_2 V_2 \quad P_1 = 1.01 \text{ atm} \quad V_1 = 5.5 \text{ L} \quad V_2 = 3.0 \text{ L}$$

$$P_2 = \frac{P_1 V_1}{V_2} = \frac{1.01 \text{ atm} \times 5.5 \text{ L}}{3.0 \text{ L}} = 1.87 \text{ atm}$$

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Table of Contents Note Taking Guide Episode 902 Chemistry

1. Understanding the eBook Note Taking Guide Episode 902 Chemistry
 - The Rise of Digital Reading Note Taking Guide Episode 902 Chemistry
 - Advantages of eBooks Over Traditional Books
2. Identifying Note Taking Guide Episode 902 Chemistry
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Note Taking Guide Episode 902 Chemistry
 - User-Friendly Interface
4. Exploring eBook Recommendations from Note Taking Guide Episode 902 Chemistry
 - Personalized Recommendations
 - Note Taking Guide Episode 902 Chemistry User Reviews and Ratings
 - Note Taking Guide Episode 902 Chemistry and Bestseller Lists
5. Accessing Note Taking Guide Episode 902 Chemistry Free and Paid eBooks
 - Note Taking Guide Episode 902 Chemistry Public Domain eBooks
 - Note Taking Guide Episode 902 Chemistry eBook Subscription Services
 - Note Taking Guide Episode 902 Chemistry Budget-Friendly Options

6. Navigating Note Taking Guide Episode 902 Chemistry eBook Formats
 - ePub, PDF, MOBI, and More
 - Note Taking Guide Episode 902 Chemistry Compatibility with Devices
 - Note Taking Guide Episode 902 Chemistry Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Note Taking Guide Episode 902 Chemistry
 - Highlighting and Note-Taking Note Taking Guide Episode 902 Chemistry
 - Interactive Elements Note Taking Guide Episode 902 Chemistry
8. Staying Engaged with Note Taking Guide Episode 902 Chemistry
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Note Taking Guide Episode 902 Chemistry
9. Balancing eBooks and Physical Books Note Taking Guide Episode 902 Chemistry
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Note Taking Guide Episode 902 Chemistry
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Note Taking Guide Episode 902 Chemistry
 - Setting Reading Goals Note Taking Guide Episode 902 Chemistry
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Note Taking Guide Episode 902 Chemistry
 - Fact-Checking eBook Content of Note Taking Guide Episode 902 Chemistry
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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