

Chemical Formula

Talking about proportion molecules and relative of atoms present in a compound

Molecular $\text{C}_{10}\text{H}_{12}\text{O}_4$ (molecular)

Formula represents the number of atoms of each element in a single molecule

Empirical $\text{C}_5\text{H}_6\text{O}_2$ (molecular)

Formula represents the simplest ratio of positive and negative ions in a compound

Atomic Mass

The average mass of atoms present in a sample of an element

Isotopes

^{12}C and ^{13}C isotopes

Same no. of protons neutrons electrons

Atomicity

^{12}C diatomic

Same no. of protons neutrons electrons

^{12}C

Note Taking Guide Episode 6chemical Formula

SJ Ball



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