

Chemical Formula

Talking about proportions molecules and relative of atoms present (e.g. hydrogen)

Molecular $\text{C}_{10}\text{H}_{22}\text{O}$ (molecular)

Formula represents the number of atoms of each in a single molecule

Empirical $\text{C}_{10}\text{H}_{22}\text{O}$ (molecular)

Formula represents the proportions of atoms of each present and negative charge formula

Atomic Mass

The proportions that atoms combine together

Mass

$\text{C}_{10}\text{H}_{22}\text{O}$ atoms

atoms proportions charge

Structure

atoms atoms

atoms proportions charge

atoms

Note Taking Guide Episode 601 Chemical Formula Key

D Kirk



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