THERMOCHEMICAL EQUATIONS AND TYPES OF ENTHALPY CHANGES

The

Thermochemical Equations

A thermochemical equation is a reaction equation accompanied by an enthalpy change (ΔH). In addition to stating the number of moles reactants and the number of moles product, it also displays the amount of heat released or absorbed in the reaction in units of kJ.

Example:

The thermochemical equation is known as:

$$H_{2(g)} + \frac{1}{2}O_{2(g)} \rightarrow H_{2}O_{00}$$
 $\Delta H = -285,85 \text{ kJ}$

This means that the formation of 1 mol H₂O hydrogen gas and oxygen gas released energy of 285,85 kJ (exothermic reaction).

Some things that must be considered in the thermochemical equation are

1. THE REACTION COEFFICIENT EXPRESSES MOLES.

Example:

$$H_{2(g)} + I_{2(g)} \rightarrow 2HI_{(g)}$$
 $\Delta H = +51,9 \text{ kJ/mol}$

(means that when 1 mol H_2 reacts with 1 mol I_2 to produce 2 mol H_2 , heat is absorbed by 51.9 kJ).

2. IF THE REACTION IS REVERSED, AH CHANGES SIGN

From reaction 1:

$$2H(g) \rightarrow H_2(g) + I_2(g)$$
 $\Delta H = -51,9 \text{ kJ/mol}$

3. IF THE REACTIONS ARE MULTIPLIED, THEN AH IS MULTIPLIED

From reaction 1:

$$\frac{1}{2}H_{2(g)} + \frac{1}{2}I_{2(g)} \rightarrow H_{1(g)}$$
 $\Delta H = +25,95 \text{ kJ/mol}$

Notetaking Guide For Thermochemical Equations

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