#### Parts of the Atom

\*Fill in the following table about the parts of the atom. (AMU stands for atomic mass unit)

PART OF THE ATOM	ELECTRIC CHARGE	LOCATION IN ATOM	ATOMIC MASS	
proton	proton	Nucleus	2000 X More than an e	
neutron	neutron	Nucleus	2000 X More than an e'	
electron	electron	Orbitals / Energy Levels	1	

\*Here's how to predict the structure of any element on the periodic table. For example, look at carbon:

6 C carbon 12.011

Error: Reference source not foundatomic number: This is the number of protons (also the number of electrons)

Error: Reference source not foundchemical symbol: This is an abbreviation for the element.

Error: Reference source not foundelement name: This is the name of the element

Error: Reference source not foundatomic mass: This number (rounded off) tells you the number of protons plus the number of neutrons. (Electrons are too small to be included.)

To find the number of protons, look at the atomic number:

Example for carbon: Atomic number: 6

Number of protons: 6

To find the number of neutrons, subtract the atomic number from the atomic mass:

Example for carbon: Atomic mass 12 (rounded) Number of neutrons 12-6=6

To find the number of electrons, look at the atomic number:

Example for carbon: Atomic number: 6 Number of electrons: 6

\*Complete the table for the elements with atomic numbers 1-10.

ELEMENT NAME	ATOMIC NUMBER	ATOMIC MASS	PROTONS	NEUTRONS	ELECTRONS
Hydrogen	1	1.00	1	1	1
Helium	2	4.00	2	2	2
Lithium	3	6.91	3	3	3
Beryllium	4	9.01	4	4	4
Boron	5	10.81	5	5	5
Carbon	6	12.01	6	6	6
Nitrogen	7	14.00	7	7	7
Oxygen	8	15.99	8	8	8

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