

c. Standard Enthalpy of Vaporization ΔH_{vap}°

→ The enthalpy change for the vaporization of 1 mole of liquid to a gas at the boiling point and standard pressure.

Example:

$$\Delta H_{vap}^\circ = \Delta H_f^\circ \quad \Delta H_{vap}^\circ = +44.0 \text{ kJ mol}^{-1}$$

d. Standard Enthalpy of Melting ΔH_{fus}°

→ The enthalpy change for melting 1 mol of a solid into a liquid at the melting point and standard pressure.

Example:

$$\Delta H_{fus}^\circ = \Delta H_{cr}^\circ \quad \Delta H_{fus}^\circ = +6.01 \text{ kJ mol}^{-1}$$

e. Standard Neutralization Enthalpy Change ΔH_{neut}°

→ The enthalpy change in the neutralization of an acid (H⁺) by a base (OH⁻) forming 1 mole of water.

Example:

$$\Delta H_{neut}^\circ = \Delta H_{acid}^\circ + \Delta H_{base}^\circ = \Delta H_{neut}^\circ = \Delta H_{cr}^\circ \quad \Delta H_{neut}^\circ = -57.3 \text{ kJ mol}^{-1}$$

f. Standard Enthalpy of Formation ΔH_{form}°

→ The enthalpy change for forming 1 mol of elemental atoms in the gas phase at standard conditions.

Example:

$$\Delta H_{form}^\circ = \Delta H_{cr}^\circ \quad \Delta H_{form}^\circ = +46.0 \text{ kJ mol}^{-1}$$

g. Standard Solution Enthalpy Change ΔH_{sol}°

→ The enthalpy change when 1 mole of a compound is converted into a solution at standard conditions.

Example:

$$\Delta H_{sol}^\circ = \Delta H_{neut}^\circ \quad \Delta H_{sol}^\circ = -57.3 \text{ kJ mol}^{-1}$$

Note Taking Guide For Thermochemical Equations

Puja Mehta



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