

# BALANCING CHEMICAL EQUATIONS

Name Grove

Rewrite and balance the equations below.

- $N_2 + H_2 \rightarrow NH_3$   $N_2 + 3H_2 \rightarrow 2NH_3$
- $KClO_3 \rightarrow KCl + O_2$   $2KClO_3 \rightarrow 2KCl + 3O_2$
- $NaCl + F_2 \rightarrow NaF + Cl_2$   $2NaCl + F_2 \rightarrow 2NaF + Cl_2$
- $H_2 + O_2 \rightarrow H_2O$   $2H_2 + O_2 \rightarrow 2H_2O$
- $AgNO_3 + MgCl_2 \rightarrow AgCl + Mg(NO_3)_2$   $2AgNO_3 + MgCl_2 \rightarrow 2AgCl + Mg(NO_3)_2$
- $AlBr_3 + K_2SO_4 \rightarrow KBr + Al_2(SO_4)_3$   $2AlBr_3 + 3K_2SO_4 \rightarrow 6KBr + Al_2(SO_4)_3$
- $CH_4 + O_2 \rightarrow CO_2 + H_2O$   $CH_4 + 2O_2 \rightarrow CO_2 + 2H_2O$
- $C_3H_8 + O_2 \rightarrow CO_2 + H_2O$   $C_3H_8 + 5O_2 \rightarrow 3CO_2 + 4H_2O$
- $C_8H_{18} + O_2 \rightarrow CO_2 + H_2O$   $C_8H_{18} + \frac{25}{2}O_2 \rightarrow 8CO_2 + 9H_2O$
- $FeCl_3 + NaOH \rightarrow Fe(OH)_3 + NaCl$   $FeCl_3 + 3NaOH \rightarrow Fe(OH)_3 + 3NaCl$
- $P + O_2 \rightarrow P_2O_5$   $4P + 5O_2 \rightarrow 2P_2O_5$
- $Na + H_2O \rightarrow NaOH + H_2$   $2Na + 2H_2O \rightarrow 2NaOH + H_2$
- $Ag_2O \rightarrow Ag + O_2$   $2Ag_2O \rightarrow 4Ag + O_2$
- $S_8 + O_2 \rightarrow SO_3$   $S_8 + 12O_2 \rightarrow 8SO_3$
- $CO_2 + H_2O \rightarrow C_6H_{12}O_6 + O_2$   $6CO_2 + 6H_2O \rightarrow C_6H_{12}O_6 + 6O_2$
- $K + MgBr_2 \rightarrow KBr + Mg$   $2K + MgBr_2 \rightarrow 2KBr + Mg$
- $HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$   $2HCl + CaCO_3 \rightarrow CaCl_2 + H_2O + CO_2$

# Physical Science If8767 Answer Key Word Equations

**JL Elias**



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a 1 only b 2 only c both 1 and 2 d neither 1 nor 2 answer d neither will be a buffer solution because buffers are something and their conjugate q which pair of solutes could be used to prepare an aqueous buffer

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