(2)

(1)

(3) [13]

#### QUESTION 2

2.1 The graph of molecular size versus the boiling point is given below.

The letters A, B and C represent the compounds CH<sub>4</sub>, C<sub>4</sub>H<sub>6</sub> and C<sub>3</sub>H<sub>8</sub> respectively.



- 2.1.1 Define the term boiling point.
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph.
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (A, B or C) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph.
  (2)
- 2.2 Consider the two molecules in the table below.

NAME OFSUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT
Ammonia	NH <sub>3</sub>	17	- 33
Phosphine	PH <sub>3</sub>	×	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES.

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**Mark Steyvers** 

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#### **Table of Contents Physical Science Department Grade Question Paper**

- 1. Understanding the eBook Physical Science Department Grade Question Paper
  - The Rise of Digital Reading Physical Science Department Grade Question Paper
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Department Grade Question Paper
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Department Grade Question Paper
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Department Grade Question Paper
  - Personalized Recommendations
  - Physical Science Department Grade Question Paper User Reviews and Ratings
  - Physical Science Department Grade Question Paper and Bestseller Lists
- 5. Accessing Physical Science Department Grade Question Paper Free and Paid eBooks
  - Physical Science Department Grade Question Paper Public Domain eBooks
  - Physical Science Department Grade Question Paper eBook Subscription Services
  - Physical Science Department Grade Question Paper Budget-Friendly Options

- 6. Navigating Physical Science Department Grade Question Paper eBook Formats
  - o ePub, PDF, MOBI, and More
  - Physical Science Department Grade Question Paper Compatibility with Devices
  - Physical Science Department Grade Question Paper Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Physical Science Department Grade Question Paper
  - Highlighting and Note-Taking Physical Science Department Grade Question Paper
  - Interactive Elements Physical Science Department Grade Question Paper
- 8. Staying Engaged with Physical Science Department Grade Question Paper
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Department Grade Question Paper
- 9. Balancing eBooks and Physical Books Physical Science Department Grade Question Paper
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Physical Science Department Grade Question Paper
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - $\circ \ Managing \ Screen \ Time$
- 11. Cultivating a Reading Routine Physical Science Department Grade Question Paper
  - Setting Reading Goals Physical Science Department Grade Question Paper
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Department Grade Question Paper
  - Fact-Checking eBook Content of Physical Science Department Grade Question Paper
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements

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