

**Chapter 10 Nuclear Chemistry**

7. How is the product isotope different from the reactant isotope in alpha decay? The product isotope has two fewer protons and two fewer neutrons than the reactant isotope.
8. Circle the letters that identify each event that takes place during beta decay.
- ☐ A proton decomposes into a neutron and an electron.
  - ☒ A neutron decomposes into a proton and an electron.
  - ☒ An electron is emitted from the nucleus.
  - ☐ A neutron is emitted from the nucleus.
9. Why are beta particles more penetrating than alpha particles? Beta particles have a smaller mass and a faster speed.
10. Is the following sentence true or false? All nuclear radiation consists of charged particles. false
11. What is a gamma ray? A gamma ray is a penetrating ray of energy emitted by an unstable nucleus.
12. How fast do gamma rays travel through space? Gamma rays travel through space at the speed of light.
13. Complete the following table about nuclear radiation.

Characteristics of Nuclear Radiation			
Radiation Type	Charge	Mass (amu)	Usually Stopped By
Alpha particle	- 1	4	Paper or clothing
Beta particle	1 -	1 1836	Aluminum sheet
Gamma ray	0	0	Several meters of concrete

**Effects of Nuclear Radiation (pages 296–297)**

14. How does nuclear radiation affect atoms? Nuclear radiation can ionize atoms.
15. Is the following sentence true or false? One potential danger of radon gas is that prolonged exposure to it can lead to lung cancer. true

**Detecting Nuclear Radiation (page 297)**

16. Name two devices that are used to detect nuclear radiation.
- Geiger counters
  - Film badges

# Physical Science Chapter 10 Nuclear Chemistry

## Wordwise Answer Key

**Christian G. Meyer**



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### **Table of Contents Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key**

1. Understanding the eBook Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
  - The Rise of Digital Reading Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
  - Personalized Recommendations
  - Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key User Reviews and Ratings
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  - Highlighting and Note-Taking Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
  - Interactive Elements Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
8. Staying Engaged with Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
9. Balancing eBooks and Physical Books Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
- Setting Reading Goals Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
- Fact-Checking eBook Content of Physical Science Chapter 10 Nuclear Chemistry Wordwise Answer Key
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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