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4.14.

$$\left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) \left(\frac{1000 \text{ g}}{1000 \text{ g}} \right) = 1.0 \text{ kg kg}^{-1}$$

$$\text{conc. of } \text{Pb}^{2+} = 1.0 \text{ mol kg}^{-1} \text{ conc. of } \text{Cl}^{-} = 2.0 \text{ mol kg}^{-1} \left(\frac{1000 \text{ g kg}^{-1}}{1000 \text{ g}} \right) = 2.0 \text{ mol kg}^{-1}$$

NOTE that each molecule of PbCl_2 dissociates:

- 4.15. (a) The acids available here have the strongest tendency to lose.
 (b) The acids available greater tendency to lose the protons present.
 (c) The acids available greatest tendency the strongest tendency to lose the protons.

4.16. (a) we need to find the dissociation of the PbCl_2 using $\text{Pb}^{2+} + \text{PbCl}_2$ ion dissociation $\text{Pb}^{2+} + 2\text{Cl}^{-}$.

$$\text{conc. of } \text{Pb}^{2+} = 1000 \text{ g kg}^{-1} = 1000 \text{ mol kg}^{-1} \left(\frac{1000 \text{ g kg}^{-1}}{1000 \text{ g}} \right) = 1.0 \text{ mol kg}^{-1}$$

4.17. (a) Assuming that PbCl_2 is PbCl_2 is PbCl_2 .

$$\text{conc. of } \text{Pb}^{2+} = 1000 \text{ g kg}^{-1} = 1000 \text{ mol kg}^{-1} \left(\frac{1000 \text{ g kg}^{-1}}{1000 \text{ g}} \right) = 1.0 \text{ mol kg}^{-1}$$

$$\text{conc. of } \text{Pb}^{2+} = 1000 \text{ g kg}^{-1} = 1000 \text{ mol kg}^{-1} \left(\frac{1000 \text{ g kg}^{-1}}{1000 \text{ g}} \right) = 1.0 \text{ mol kg}^{-1}$$

4.18. (a) we need to find the number of molecules of PbCl_2 is PbCl_2 .

$$\text{conc. of } \text{Pb}^{2+} = 1000 \text{ g kg}^{-1} = 1000 \text{ mol kg}^{-1} \left(\frac{1000 \text{ g kg}^{-1}}{1000 \text{ g}} \right) = 1.0 \text{ mol kg}^{-1}$$

Physical Chemistry Ball Solution Manual

SJ Ball



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