

Chapter 6: Chemical Bonds Review Answer Key

- What is the most stable group on the Periodic Table? Why is this group stable?
Noble Gases (Group 18) because they have 8 valence electrons which means their outer energy level is full.
- Name the most reactive group(s) on the Periodic Table? Why are they reactive?
Alkali Metals (Group 1) & Halogens (Group 17): Alkali Metals want to lose one valence electron in order to become stable and Halogens want to gain one valence electron to be stable.

- In the boxes below draw the Electron dot diagrams for:

Sodium	Calcium	Boron	Chlorine	Krypton
Na ·	Ca · ·	· B · ·	· · · : Cl · · ·	· · · : Kr : · · ·

- Name the three (3) types of chemical bonds:
Ionic bonds Covalent bonds Polyatomic
- Define ionic bond:
The force that holds cations and anions together.
- Define covalent bond:
A chemical bond in which two atoms share a pair of valence electrons.
- Define polyatomic ion:
A covalently bonded group of atoms that has a positive or negative charge and acts as a unit.
- In an ionic bond, electrons are transferred.
- Who are the participants in an ionic bond? **Metals and nonmetals**
- An ion with a positive charge is a cation; an ion with a negative charge is an anion.
- In a covalent bond, electrons are shared.
- Who are the participants in a covalent bond? **Two nonmetals**
- Identify whether the following compounds are covalent (two nonmetals), ionic (a nonmetal and a metal), or polyatomic (both).
 - CaCl₂ IONIC (transferred)**
 - H₂O COVALENT (shared)**
 - BaSO₄ POLYATOMIC (both)**
 - CH₄ COVALENT (shared)**
 - LiBr IONIC (transferred)**
 - NH₄Cl POLYATOMIC (both)**
 - KI IONIC (transferred)**

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DP Hallahan



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