

1. Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in atomic radius as you go across a period is DECREASING. Therefore, the element on the far left (K) is the largest, and the element on the far right (Br) is the smallest.*

2. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in ionization energy as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest ionization energy, and the element on the far right (Br) has the highest ionization energy.*

3. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity:

Cu   **K**   Ni   **Br**

- a. Explain why you made these choices: *All of the elements are in the same period. The trend in electronegativity as you go across a period is INCREASING. Therefore, the element on the far left (K) has the lowest electronegativity, and the element on the far right (Br) has the highest electronegativity.*

4. For each of the following groups: Circle the element with the largest atomic radius and put a square around the element with the smallest atomic radius:

- 5.
- O   C   **Be**   **Ne**   Same Period
  - Na   Rb   **Fr**   **H**   Same Group
  - Pb**   **C**   Sn   Si   Same Group
  - Au   W   S   **Fr**   **Ne**   Zn   Challenge

6. For each of the following groups: Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy:

- O   C   **Be**   **Ne**   Same Period
- Na   Rb   **Fr**   **H**   Same Group
- Pb**   **C**   Sn   Si   Same Group
- Au   W   S   **Fr**   **Ne**   Zn   Challenge

# Periodic Trends Practice Problems Answers Key Bing

**Rachel S Tattersall**



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**Scientific and Technical Aerospace Reports**, 1984 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA

Scientific and Technical Information Database      *Homenatge professor Josep M. Costa. Trends in electrochemistry and corrosion at the beginning of the 21st century* Pere-Lluís Cabot Julià, 2004-04-13 Este libro está dedicado al Profesor Josep M Costa en ocasión de su 70 aniversario. Reúne un total de 73 artículos y revisiones originales tanto científicas como tecnológicas escritas en español e inglés por unos 250 investigadores de todo el mundo y que son exponentes representativos de la investigación internacional en materias de gran interés en la Electroquímica y la Corrosión de principios de este siglo XXI. El libro se ha estructurado en dos grandes secciones. La primera sección correspondiente a la Electroquímica consta de 33 trabajos distribuidos en 5 capítulos dedicados a los campos de Electroquímica Molecular, Electrodeposición, Electroodos Modificados, Descontaminación Electroquímica y Sensores y Electroanálisis. La segunda sección relativa a la Corrosión comprende 40 trabajos que se agrupan en otros 5 capítulos que versan sobre Corrosión en Ambientes Corrosivos, Selección de Protección contra la Corrosión y Monitorización, Recubrimientos, Nuevos Materiales y Tratamientos y Educación en la Corrosión. This book is dedicated to Professor Josep M Costa in occasion of his 70th birthday. It collects a total number of 73 original articles and reviews both scientific and technologic written in English and Spanish by about 250 researchers of all around the world who are representative exponents of the international research in topics of great interest in Electrochemistry and Corrosion at the beginning of the 21st Century. The book has been structured in two large sections. The first section corresponds to Electrochemistry and includes 33 articles distributed into five chapters related to the fields of Molecular Electrochemistry, Electrodeposition, Modified Electrodes, Electrochemical Depollution and Sensors and Electroanalysis. The second section is related to Corrosion and contains 40 articles gathered into other five chapters devoted to Corrosion in Selected Environments, Corrosion Protection and Monitoring, Coatings, New Materials and Treatments and Corrosion Education.

**Shriver and Atkins' Inorganic Chemistry** Peter Atkins, 2010 Inorganic Chemistry fifth edition represents an integral part of a student's chemistry education. Basic chemical principles are set out clearly in Foundations and are fully developed throughout the text, culminating in the cutting edge research topics of the Frontiers which illustrate the dynamic nature of inorganic chemistry.

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