

# Percent Yield Practice

3 pages + 3-page Answer KEY + Periodic Table

Chemistry

## PERCENT YIELD

Show all work and circle or highlight your final answer.

$$\text{percent yield (\%)} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

- 1 16.8 grams of sodium reacts with excess chlorine gas.



The theoretical yield of sodium chloride is 25.5 grams. The actual yield was 23.5 grams. Calculate the percent yield of sodium chloride.

Name \_\_\_\_\_

Date \_\_\_\_\_

- 2 Hydrogen gas ( $\text{H}_2$ ) reacts with oxygen gas ( $\text{O}_2$ ) to produce water ( $\text{H}_2\text{O}$ ). 5.8 grams of hydrogen reacts with excess oxygen gas.



What is the theoretical yield?

- 3 8.8 grams of carbon reacts with excess oxygen gas.



The theoretical yield of carbon dioxide is 25.5 g. Calculate the percent yield of carbon dioxide.

- 4 1.8 grams of zinc reacts with excess sulfur.



The theoretical yield of zinc sulfide is 3.5 grams. Calculate the percent yield of zinc sulfide.

- 5 The actual yield was 28.8 grams. Calculate the percent yield.

- 6 Magnesium ( $\text{Mg}$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce magnesium chloride ( $\text{MgCl}_2$ ) and hydrogen gas ( $\text{H}_2$ ). 10.0 grams of magnesium reacts with excess hydrochloric acid.



What is the theoretical yield?

- 7 1.0 grams of phosphorus reacts with excess chlorine gas.



The theoretical yield of phosphorus trichloride is 3.5 grams. Calculate the percent yield of phosphorus trichloride.

- 8 The actual yield was 40.2 grams. Calculate the percent yield.

Name \_\_\_\_\_

Date \_\_\_\_\_

- 9 Calcium carbonate ( $\text{CaCO}_3$ ) reacts with hydrochloric acid ( $\text{HCl}$ ) to produce calcium chloride ( $\text{CaCl}_2$ ), carbon dioxide ( $\text{CO}_2$ ), and water ( $\text{H}_2\text{O}$ ). 15.0 grams of calcium carbonate reacts with excess hydrochloric acid.



What is the theoretical yield of calcium chloride?

$$\begin{array}{r} 15.0 \text{ g CaCO}_3 \times \frac{1 \text{ mol CaCO}_3}{100.09 \text{ g CaCO}_3} \times \frac{1 \text{ mol CaCl}_2}{1 \text{ mol CaCO}_3} \times \frac{146.97 \text{ g CaCl}_2}{1 \text{ mol CaCl}_2} = 13.9 \text{ g CaCl}_2 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCl}_2 \\ \text{Ca} = 40.08 \\ \text{Cl} = 35.45 \\ 2 \times 35.45 = 70.90 \\ 40.08 + 70.90 = 110.98 \text{ g} \end{array}$$

10 The actual yield was 13.9 grams. Calculate the percent yield of calcium chloride.

$$\begin{array}{r} 13.9 \text{ g CaCl}_2 \times \frac{1 \text{ mol CaCl}_2}{146.97 \text{ g CaCl}_2} \times \frac{1 \text{ mol CaCO}_3}{1 \text{ mol CaCl}_2} \times \frac{100.09 \text{ g CaCO}_3}{1 \text{ mol CaCO}_3} = 9.5 \text{ g CaCO}_3 \end{array}$$

$$\begin{array}{r} \text{Molar Mass CaCO}_3 \\ \text{Ca} = 40.08 \\ \text{C} = 12.01 \\ 3 \times 16.00 = 48.00 \\ 40.08 + 12.01 + 48.00 = 100.09 \text{ g} \end{array}$$

- 11 The actual yield was 25.5 grams. Calculate the percent yield of calcium chloride.

$$\text{percent yield (\%)} = \frac{25.5 \text{ g}}{100.09 \text{ g}} \times 100 = 25.5\%$$

- 12 Potassium ( $\text{K}$ ) reacts with iodine ( $\text{I}_2$ ) to produce potassium iodide ( $\text{KI}$ ). 10.0 grams of potassium reacts with excess iodine.



What is the theoretical yield?

$$\begin{array}{r} 10.0 \text{ g K} \times \frac{1 \text{ mol K}}{39.10 \text{ g K}} \times \frac{2 \text{ mol KI}}{2 \text{ mol K}} \times \frac{166.10 \text{ g KI}}{1 \text{ mol KI}} = 166.10 \text{ g KI} \end{array}$$

$$\begin{array}{r} \text{Molar Mass KI} \\ \text{K} = 39.10 \\ \text{I} = 126.90 \\ 39.10 + 126.90 = 166.00 \text{ g} \end{array}$$

YIELD

ANSWER KEY INCLUDED

## Percent Yield Example Answer Key

**M Tight**



**Percent Yield Example Answer Key:**

## Whispering the Techniques of Language: An Mental Journey through **Percent Yield Example Answer Key**

In a digitally-driven earth where screens reign supreme and instant interaction drowns out the subtleties of language, the profound techniques and psychological nuances hidden within words usually get unheard. However, located within the pages of **Percent Yield Example Answer Key** a interesting literary prize sporting with organic thoughts, lies a fantastic journey waiting to be undertaken. Penned by a skilled wordsmith, that wonderful opus encourages viewers on an introspective trip, delicately unraveling the veiled truths and profound influence resonating within the very cloth of each word. Within the psychological depths with this emotional review, we can embark upon a honest exploration of the book is key subjects, dissect their captivating writing fashion, and yield to the powerful resonance it evokes strong within the recesses of readers hearts.

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