

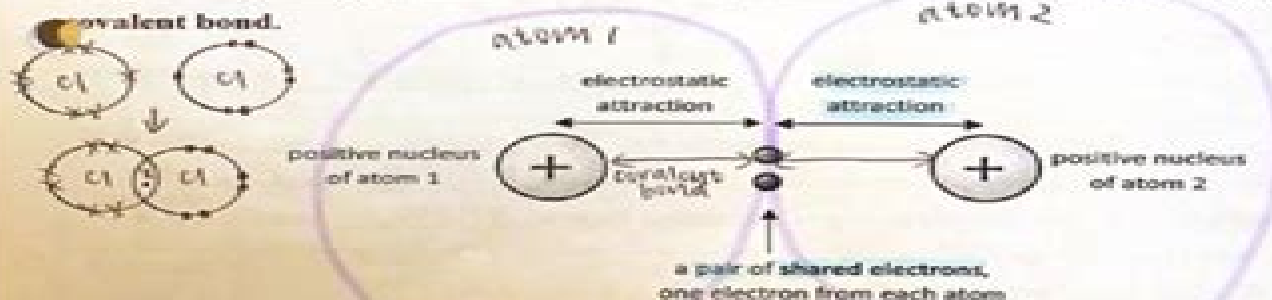
Chapter 8: Covalent Bonding

Students should:

- 1.14 know what is meant by the term molecule
- 1.44 know that a covalent bond is formed between atoms by the sharing of a pair of electrons
- 1.45 understand covalent bonds in terms of electrostatic attractions
- 1.46 understand how to use dot-and-cross diagrams to represent covalent bonds in:
 - diatomic molecules, including hydrogen, oxygen, nitrogen, halogens and hydrogen halides
 - inorganic molecules including water, ammonia and carbon dioxide
 - organic molecules containing up to two carbon atoms, including methane, ethane, ethene and those containing halogen atoms
- 1.47 explain why substances with simple molecular structures are gases or liquids, or solids with low melting and boiling points
the term intermolecular forces of attraction can be used to represent all forces between molecules
- 1.48 explain why the melting and boiling points of substances with simple molecular structures increase, in general, with increasing relative molecular mass
- 1.49 explain why substances with giant covalent structures are solids with high melting and boiling points
- 1.50 explain how the structures of diamond, graphite and C_{60} fullerene influence their physical properties, including electrical conductivity and hardness
- 1.51 know that covalent compounds do not usually conduct electricity

8.1 Covalent bonding

A covalent bond is resulted when non-metal atoms combine with non-metal atoms by sharing of outermost shell electrons. Strong, directional electrostatic attraction is formed between the (negatively charged) shared electrons and the (positively charged) nuclei of the two bonded atoms. Such attraction is called a



The bonded atoms can then attain an octet or duplet of electrons and become stable (i.e. each atom attains the stable electronic structure of its nearest noble gas).

Pearson Education Chapter 8 Covalent Bonding Review Answers

Yan Bai



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Table of Contents Pearson Education Chapter 8 Covalent Bonding Review Answers

1. Understanding the eBook Pearson Education Chapter 8 Covalent Bonding Review Answers
 - The Rise of Digital Reading Pearson Education Chapter 8 Covalent Bonding Review Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Education Chapter 8 Covalent Bonding Review Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Education Chapter 8 Covalent Bonding Review Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Education Chapter 8 Covalent Bonding Review Answers
 - Personalized Recommendations
 - Pearson Education Chapter 8 Covalent Bonding Review Answers User Reviews and Ratings
 - Pearson Education Chapter 8 Covalent Bonding Review Answers and Bestseller Lists
5. Accessing Pearson Education Chapter 8 Covalent Bonding Review Answers Free and Paid eBooks
 - Pearson Education Chapter 8 Covalent Bonding Review Answers Public Domain eBooks
 - Pearson Education Chapter 8 Covalent Bonding Review Answers eBook Subscription Services
 - Pearson Education Chapter 8 Covalent Bonding Review Answers Budget-Friendly Options
6. Navigating Pearson Education Chapter 8 Covalent Bonding Review Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Pearson Education Chapter 8 Covalent Bonding Review Answers Compatibility with Devices
 - Pearson Education Chapter 8 Covalent Bonding Review Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Pearson Education Chapter 8 Covalent Bonding Review Answers
 - Highlighting and Note-Taking Pearson Education Chapter 8 Covalent Bonding Review Answers
 - Interactive Elements Pearson Education Chapter 8 Covalent Bonding Review Answers
8. Staying Engaged with Pearson Education Chapter 8 Covalent Bonding Review Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Education Chapter 8 Covalent Bonding Review Answers
9. Balancing eBooks and Physical Books Pearson Education Chapter 8 Covalent Bonding Review Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Education Chapter 8 Covalent Bonding Review Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pearson Education Chapter 8 Covalent Bonding Review Answers
- Setting Reading Goals Pearson Education Chapter 8 Covalent Bonding Review Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pearson Education Chapter 8 Covalent Bonding Review Answers
- Fact-Checking eBook Content of Pearson Education Chapter 8 Covalent Bonding Review Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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