

Chemistry 12 Worksheet 2-3

[KEY]

Calculations Involving the Equilibrium Constant K_{eq}

1. Given the equilibrium equation below:



If, at equilibrium, the concentrations are as follows:

$$[A_2] = 3.45 \text{ M}, \quad [B_2] = 5.67 \text{ M} \quad \text{and} \quad [AB] = 0.67 \text{ M} \quad (2 \text{ SD})$$

- a) Write the expression for the equilibrium constant,
- K_{eq}

$$K_{eq} = \frac{[AB]^2}{[A_2][B_2]}$$

- b) Find the value of the equilibrium constant,
- K_{eq}
- at the temperature that the experiment was done.

$$K_{eq} = \frac{(0.67)^2}{(3.45)(5.67)} = 0.023$$

Answer $K_{eq} = 0.023$

2. Given the equilibrium equation:

at a temperature of 50°C , it is found that when equilibrium is reached that:

$$[X_2] = 0.37 \text{ M}, \quad [Y_2] = 0.53 \text{ M} \quad \text{and} \quad [XY_3] = 0.090 \text{ M} \quad (\text{All } 2 \text{ SD})$$

- a) Write the equilibrium constant expression (
- K_{eq}
-)

$$K_{eq} = \frac{[XY_3]^2}{[X_2][Y_2]^3}$$

- b) Calculate the value of
- K_{eq}
- at
- 50°C
- .

$$K_{eq} = \frac{(0.090)^2}{(0.37)(0.53)^3} = 0.15$$

Answer $K_{eq} = 0.15$

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Stoichiometry

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