

6

THE PERIODIC TABLE

Chapter Quiz

Fill in the word(s) that will make each statement true.

- | | |
|--|--------------|
| 1. In the modern periodic table, when elements are arranged according to their atomic <u>1</u> , there is a periodic repetition of properties. | 1. _____ 6.1 |
| 2. There are <u>2</u> periods in the periodic table. | 2. _____ 6.1 |
| 3. The elements in any <u>3</u> in the periodic table have similar physical and chemical properties. | 3. _____ 6.1 |
| 4. Oxygen and sulfur, Group 6A elements, have <u>4</u> electrons in their highest occupied energy level. | 4. _____ 6.2 |
| 5. For the inner transition elements, electrons are added to an <i>f</i> sublevel with a principal energy level that is <u>5</u> than the period number. | 5. _____ 6.2 |

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- | | |
|---|-----|
| _____ 6. The atomic radius of an element in period 3 is larger than the atomic radius of an element in period 2. | 6.3 |
| _____ 7. For Group 3A elements, there is a relatively small increase in ionization energy between the second and third ionization energies. | 6.3 |
| _____ 8. Anions are smaller than the neutral atoms from which they are formed. | 6.3 |
| _____ 9. Atoms with low electronegativity values tend to form positive ions. | 6.3 |
| _____ 10. As a group, alkali metals have the highest electronegativities. | 6.3 |

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M Walker



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1. Understanding the eBook Pearson Chemistry The Periodic Table Answers
 - The Rise of Digital Reading Pearson Chemistry The Periodic Table Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry The Periodic Table Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry The Periodic Table Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry The Periodic Table Answers
 - Personalized Recommendations
 - Pearson Chemistry The Periodic Table Answers User Reviews and Ratings

- Pearson Chemistry The Periodic Table Answers and Bestseller Lists
- 5. Accessing Pearson Chemistry The Periodic Table Answers Free and Paid eBooks
 - Pearson Chemistry The Periodic Table Answers Public Domain eBooks
 - Pearson Chemistry The Periodic Table Answers eBook Subscription Services
 - Pearson Chemistry The Periodic Table Answers Budget-Friendly Options
- 6. Navigating Pearson Chemistry The Periodic Table Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pearson Chemistry The Periodic Table Answers Compatibility with Devices
 - Pearson Chemistry The Periodic Table Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pearson Chemistry The Periodic Table Answers
 - Highlighting and Note-Taking Pearson Chemistry The Periodic Table Answers
 - Interactive Elements Pearson Chemistry The Periodic Table Answers
- 8. Staying Engaged with Pearson Chemistry The Periodic Table Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry The Periodic Table Answers
- 9. Balancing eBooks and Physical Books Pearson Chemistry The Periodic Table Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry The Periodic Table Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Chemistry The Periodic Table Answers
 - Setting Reading Goals Pearson Chemistry The Periodic Table Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry The Periodic Table Answers
 - Fact-Checking eBook Content of Pearson Chemistry The Periodic Table Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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