BITTHEY.			

# Chemistry 30

# Thermochemistry Review

- 1. Which of the following statements is true for an endothermic reaction?
  - A. ΔH<sub>e</sub> for the reaction is negative.
  - B. Energy is a product in the chemical equation.
  - C. The temperature of the surroundings increases.
  - D. The potential energy of the products is greater than the potential energy of the reactants.

Use the following information to answer the next question.

$$4 \; Q(s) \; + \; Z_2(g) \; + \; 100 \; kJ \; \rightarrow \; 2 \; Q_2 Z(g)$$

- The ΔH<sub>r</sub> value for the reverse reaction would be
  - A. +100 kJ
  - B. +50.0 kJ
  - C. -50.0 kJ
  - D. -100 kJ

#### Numerical Response

Benzene is a gasoline additive. The heat of formation for 1.00 mol of C<sub>6</sub>H<sub>6</sub>(1) is kJ.

(Record your three-digit answer in the numerical-response section on the answer sheet)

#### Numerical Response

An outdoor mercury thermometer contains 1.02 g of mercury. On a winter day, the thermometer records a change in temperature from −38.0°C to −13.0°C. The mercury absorbed \_\_\_\_\_\_ J of energy. The heat capacity of mercury is 0.140 J/g°C.

(Record your three-digit answer in the numerical-response section on the answer sheet)

# **Pearson Chemistry Section Review Answers Thermochemistry**

**M** Carnoy

**Pearson Chemistry Section Review Answers Thermochemistry:** 

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