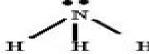
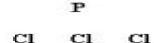
7-2 Review and Reinforcement (continued)

Write Lewis structures for each of the following molecules. Indicate the bonds with either or dashes.

15. NHL.





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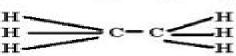
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17. C.H. н H



CI

Answer the following questions as directed.

21. Explain why the molecule SF4 is an exception to the octet rule.

In this molecule, 10 electrons surround the sulfur atom. not 8. the sulfur atom holds the additional electrons in its empty 3 d

22. Explain the relationship between electronegativity and polarity.

An atoms e(-) is a number that indicates how strongly it attracts the electrons ina chemical bond with another atom. thus the electronegativity difference between the two atoms in a bond indicates the bond's polarity, or charge shift. The greater the electronegativity, the greater the polarity of the bond

23. Compare and contrast single, double, and triple covalent bonds.

these bonds are all covalent bonds, single bonds involve one shared pair, double bonds involve two shared pares, and triple bonds involve three shared pairs.

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Emilie Sanchez

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