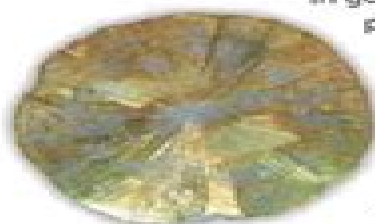


7.1 Ions

Connecting to Your World

Pyrrite (FeS_2), a common mineral that emits sparks when struck against steel, is often mistaken for gold—hence its nickname, “fool’s gold.” Although certainly not worth its weight in gold, pyrite can be used as a source of sulfur in the production of sulfuric acid, a common industrial chemical. Pyrite is an example of a crystalline solid. In crystalline solids, the component particles of the substance are arranged in an orderly, repeating fashion. In this chapter, you will learn about crystalline solids composed of ions that are bonded together. But first you need to understand how ions form from neutral atoms.



Valence Electrons

Mendeleev used similarities in the properties of elements to organize his periodic table. Scientists later learned that all of the elements within each group of the periodic table behave similarly because they have the same number of valence electrons. **Valence electrons** are the electrons in the highest occupied energy level of an element's atoms. The number of valence electrons largely determines the chemical properties of an element.

The number of valence electrons is related to the group numbers in the periodic table.  To find the number of valence electrons in an atom of a representative element, simply look at its group number. For example, the elements of Group 1A (hydrogen, lithium, sodium, potassium, and so forth) all have one valence electron, corresponding to the 1 in 1A. Carbon and silicon, in Group 4A, have four valence electrons. Nitrogen and phosphorus, in Group 5A, have five valence electrons; and oxygen and sulfur, in Group 6A, have six. The noble gases (Group 8A) are the only exceptions to the group-number rule: Helium has two valence electrons, and all of the other noble gases have eight. Figure 7.1 shows some applications of Group 4A elements.

Guide for Reading

Key Concepts

- How do you find the number of valence electrons in an atom of a representative element?
- Atoms of which elements tend to gain electrons? Atoms of which elements tend to lose electrons?
- How are cations formed?
- How are anions formed?

Vocabulary

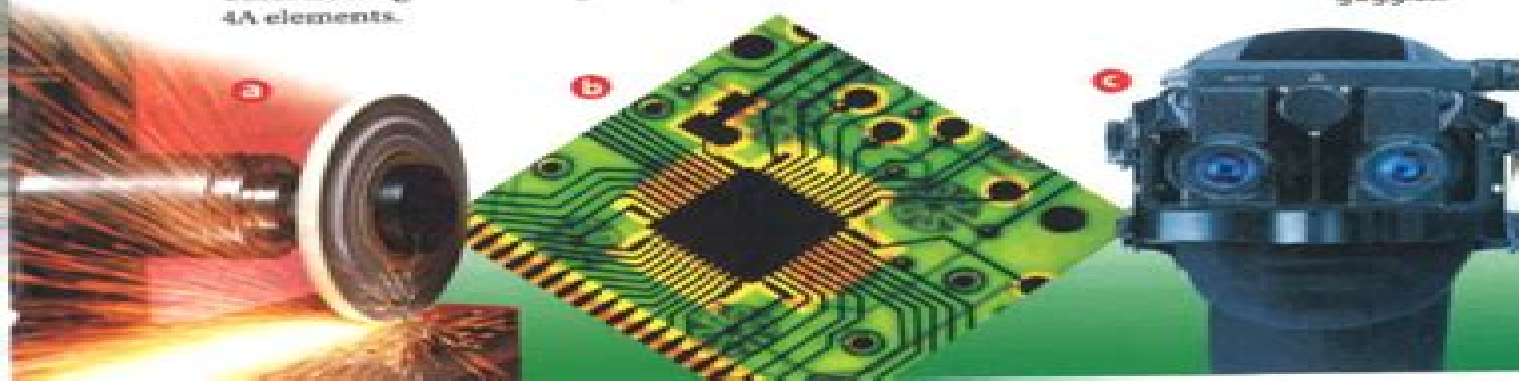
valence electrons
electron dot structures
octet rule
halide ions

Reading Strategy

Summarizing Write a one-paragraph summary of how the octet rule applies to the formation of ions.

Figure 7.1 Group 4A elements include carbon, silicon, and germanium.

 This saw blade contains carbon in the form of diamond.  Silicon is used in the manufacture of microchips.  Germanium is one of the materials used to make thermoscanning goggles.



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