

CHEM 105-Fundamentals of Chemistry

Problem Set 6-Thermodynamics

- Predict the sign (+ or -) of ΔS for the following processes:
 - $\text{Br}_2 (\text{g}) \rightarrow 2\text{Br} (\text{g})$
 - $\text{H}_2\text{O} (\text{l}) \rightarrow \text{H}_2\text{O} (\text{g})$
 - Sugar (s) $\rightarrow$ Sugar (aq)
 - $2\text{SO}_2 (\text{g}) + \text{O}_2 (\text{g}) \rightarrow 2\text{SO}_3 (\text{g})$
 - Cooling of water from 70 °C to 14 °C
- Calculate ΔS when 1 g of $\text{H}_2\text{O} (\text{l})$ is reversibly evaporated at 100 °C. ($\Delta H_{\text{vap}}=540$ cal/g)
- Some asymmetrical compounds like CO may not have third law of thermodynamics (The entropy of a pure perfect crystal at 0 K is zero). Can you explain why?
- Boiling point of propane is -22 °C at 1 atm. Calculate ΔH_{vap} for propane using Trouton's rule.
- "Dissociation of a diatomic gas is favored by high temperature". Justify this statement by using thermodynamics. Use your knowledge of bond energy and entropy.
- A traveller goes from California to Colorado. Which of the followings are state functions?
 - distance traveled
 - latitude change
 - altitude change
 - gasoline consumed
 - elapsed time
 - longitude change
- Given the following information: $\Delta H_f^\circ (\text{H}_2\text{O})(\text{g}) = -242 \text{ kJ/mol}$
 $\Delta H_f^\circ (\text{CH}_4)(\text{g}) = -74.9 \text{ kJ/mol}$
the heat of combustion of $\text{CH}_4 (\text{g})$ to $\text{CO}_2 (\text{g})$ and $\text{H}_2\text{O} (\text{g})$ is -803 kJ/mol.
 - Calculate $\Delta H_{\text{rxn}}^\circ$ for the reaction: $\text{C} (\text{s}) + 2\text{H}_2\text{O} (\text{g}) \rightarrow \text{CO}_2 (\text{g}) + 2\text{H}_2 (\text{g})$
 - Calculate ΔU° for the same reaction
- Use thermodynamic data at 298 K to decide in which direction the below reaction is spontaneous when the partial pressures of $\text{SO}_2 (\text{g})$, $\text{O}_2 (\text{g})$ and $\text{SO}_3 (\text{g})$ are 1×10^{-4} , 0.20 and 0.10 atm, respectively.
$$2\text{SO}_2 (\text{g}) + \text{O}_2 (\text{g}) \rightleftharpoons 2\text{SO}_3 (\text{g})$$
$$\Delta G_f^\circ (\text{SO}_2)(\text{g}) = -300.2 \text{ kJ/mol}$$
$$\Delta G_f^\circ (\text{SO}_3)(\text{g}) = -371.1 \text{ kJ/mol}$$

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