

6**Chemical Quantities**

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A. Completion

This activity requires students to check prior knowledge of the terms and their understanding of the concepts introduced in this chapter. Each blank can be completed with a unit, three phrases, or numbers.

Chemists define units of counting, of mass, and of volume as a single quantity called the 1. Chemists measure the "amount of substance." One way to measure the amount of a substance is to count the number of 2 of that substance. The amount of representative particles for a sample of a substance is called 3. This number is 4. For example, a mole of nitrogen contains $6.02 \times 10^{23} 5 of nitrogen.$

Chemists define gram atomic mass as the weight of grams of an element whose mass is equal to the 6 atomic mass. This is the mass of 1.0 g of carbon, equal to the atomic number of carbon as 7 g of oxygen atoms.

To find the amount in moles of a substance represented, we will require the 8 of the above-stating equal analogy. When the substance the unit mass has mass, we divide the 9 of the compound. If we divide a compound that is not an appropriate quantity, because the representative particles found in each compound is the formula unit, the compound's mole is called as 10. The formula mass formula mass and 11 can be used to refer to the mass of a mole of any substance or compound.

The gram formula mass is a mass of a substance composed of a molecule or a unit of a compound or an ion. The definition that mass of each unit is a compound or an ion, for example, we can find the mass formula as 12. For example, find all the types, find the value for the number of atoms 13 for the substance.

1. mass2. representative particles3. Avogadro's Number4. 6.02×10^{23} 5. molecules6. atomic mass7. 128. gram atomic masses9. gram molecular mass10. gram formula mass11. Molar mass12. conversion factor13. chemical formula

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