13 These calculations have used $L = 6.02 \times 10^{23}$
(a) 7.22×10^{23} (b) 3.01×10^{23}
(c) 1.20×10^{23}

14 0.53 mol H

15 0.250 mol

16 (a) 262.87 g mol^{-1} (b) 176.14 g mol^{-1}
(c) 164.10 g mol^{-1} (d) 248.22 g mol^{-1}

17 189.1 g

18 1.5 mol

19 0.0074 mol Cl^-

20 1.83×10^{23} C atoms

21 171 g (integer value because no calculator)

22 10.0 g H_2O

23 2.0 mol N_2 > 3.0 mol NH_3 > 25.0 mol H_2 > 1.0 mol N_2H_4

24 (a) CH (b) CH_2O
(c) $C_{12}H_{22}O_{11}$ (d) C_2H_5
(e) C_2H_6 (f) CH_2O

25 $Na_2S_2O_3$

26 $CoSO_4.7H_2O$

27 $C_{10}H_{22}N$

28 NH_3

29 6.94 Li

30 CdS

31 empirical formula CH; molecular formula C_2H_2

32 empirical formula H_2PO_3; molecular formula $H_4P_2O_6$

33 $C_{10}H_{12}N_2P_2O_{12}$, for both empirical and molecular formulas

34 C_2H_6O

35 Let y = mass of chalk in grams.

moles of chalk used = $\frac{\text{mass used}}{M_r(CaCO_3)} \times$

$y \frac{y \text{ g}}{100.09 \text{ g mol}^{-1}}$

This is the same as the number of moles of carbon atoms used.

Therefore the number of carbon atoms used = moles of chalk × (6.02 × 1023 mol−1) = $\frac{6.02 \times 10^{23}\ y}{100.09}$

36 (a) 2.50 mol (b) 5.63 mol
(c) 665.5 g

37 (a) $2C_4H_{10} + 13O_2 \rightarrow 8CO_2 + 10H_2O$
(b) 1.59 g

38 4.355 kg

39 (a) $CaCO_3 \rightarrow CaO + CO_2$
(b) 92.8%
(c) $CaCO_3$ is the only source of CO_2; all the $CaCO_3$ undergoes complete decomposition; all CO_2 released is captured; heating does not cause any change in the mass of the other minerals present.

40 (a) 85.2 g (b) 1.3 g H_2

41 5.23 g $C_2H_4Cl_2$

42 253.2 g theoretical $CaSO_4$; 77.9%

43 3.16 g ester

44 107 g of C_2H_5 needed

45 (a) 2.40 mol (b) 0.0110 mol
(c) 44 mol

46 (a) 35.65 dm^3 (b) 5.7 dm^3

47 0.652 dm^3

48 0.138 mol Br_2 and 0.156 mol Cl_2, so more molecules of Cl_2

49 0.113 dm^3

50 0.28 dm^3

51 90 kPa

52 16 °C

53 3.0 dm^3

54 2.8 dm^3

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AW Rasmussen

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