

Chapter 17 REVIEW

1. Given a solution of ammonium chloride. What additional reagent or reagents are needed to prepare a buffer from the ammonium chloride solution?

Since NH_4Cl is a salt of a weak base, the weak base NH_3 is needed.

- a) Explain how **this** buffer solution resists a change in pH when (show the reaction that occurs with explanation of how it buffers the solution):
- Moderate amounts of strong acid are added.



- Moderate amounts of strong base are added.



2. Describe the solubility of CaCO_3 in each of the following solutions compared to its solubility in pure water. Show all applicable reactions. Give a brief explanation:

- a) In 0.10 M NaCl $\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$ Neither of these ions are a common ion so there will be no effect on the solubility of CaCO_3 .
- b) In 0.10 M Na_2CO_3 $\text{Na}_2\text{CO}_3 \rightarrow 2\text{Na}^+ + \text{CO}_3^{2-}$ since this salt has carbonate which is a common ion it will affect the solubility of CaCO_3 . $\text{CaCO}_3(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + \text{CO}_3^{2-}(\text{aq})$ so when the concentration of carbonate is increased it will shift the reaction to the left and decrease the solubility.
- c) In 0.10 M HCl $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$ H^+ will react with CO_3^{2-} to make carbonic acid which immediately breaks down into water and carbon dioxide. $2\text{H}^+ + \text{CO}_3^{2-} \rightarrow \text{H}_2\text{O} + \text{CO}_2$ since the carbonate ion is reacted with the H^+ it reduces the concentration of carbonate in the solution and causes the reaction to shift to the right increasing the solubility.

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