

Name _____

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 1) At equilibrium, _____.
 - A) the rates of the forward and reverse reactions are equal
 - B) the rate constants of the forward and reverse reactions are equal
 - C) all chemical reactions have ceased
 - D) the value of the equilibrium constant is 1
 - E) the limiting reagent has been consumed
- 2) Which one of the following will change the value of an equilibrium constant?
 - A) adding other substances that do not react with any of the species involved in the equilibrium
 - B) varying the initial concentrations of reactants
 - C) changing temperature
 - D) varying the initial concentrations of products
 - E) changing the volume of the reaction vessel

- 3) The value of K_{eq} for the following reaction is 0.25:



The value of K_{eq} at the same temperature for the reaction below is _____.



- 4) The value of K_{eq} for the equilibrium

$$\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$$
 is 794 at 25°C. At this temperature, what is the value of K_{eq} for the equilibrium below?

$$\text{HI}(\text{g}) \rightleftharpoons 1/2 \text{H}_2(\text{g}) + 1/2 \text{I}_2(\text{g})$$

- A) 0.062 B) 16 C) 0.25 D) 0.50 E) 0.12
- A) 0.035 B) 0.0013 C) 28 D) 397 E) 1588

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SA Adler



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