



# COVALENT BONDING

## Practice Problems

Do your answers to these very practice problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following as ionic or molecular.

- a.  $\text{NaCl}$  (ionic)      b.  $\text{CH}_4$  (molecular)      c.  $\text{H}_2\text{O}$  (molecular)  
 d.  $\text{KNO}_3$  (ionic)      e.  $\text{H}_2\text{SO}_4$  (molecular)

2. Which of the following are molecular compounds?

- a.  $\text{H}_2\text{O}$       b.  $\text{CO}_2$       c.  $\text{H}_2\text{O}_2$   
 d.  $\text{NaCl}$       e.  $\text{H}_2\text{SO}_4$       f.  $\text{H}_2\text{O}_2$

3. Which of the following are molecular compounds? Write the molecular formula.

a.  $\text{Carbon dioxide}$  is a  $\text{CO}_2$  molecular compound. The molecular formula is  $\text{CO}_2$ .  
 b.  $\text{Sulfur dioxide}$  is a  $\text{SO}_2$  molecular compound. The molecular formula is  $\text{SO}_2$ .  
 c.  $\text{Hydrogen peroxide}$  is a  $\text{H}_2\text{O}_2$  molecular compound. The molecular formula is  $\text{H}_2\text{O}_2$ .

d.  $\text{Water}$  is a  $\text{H}_2\text{O}$  molecular compound. The molecular formula is  $\text{H}_2\text{O}$ .  
 e.  $\text{Hydrogen sulfide}$  is a  $\text{H}_2\text{S}$  molecular compound. The molecular formula is  $\text{H}_2\text{S}$ .

### SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride,  $\text{HF}$ .



2. Draw the Lewis structure for methane,  $\text{CH}_4$ .



3. Draw the Lewis structure for carbon dioxide,  $\text{CO}_2$ .



4. Draw the Lewis structure for acetylene,  $\text{C}_2\text{H}_2$ .

5. Draw the Lewis structure for hydrogen cyanide,  $\text{HCN}$ . Is this a molecular compound or ionic?



6. Draw the Lewis structure for hydrogen cyanide,  $\text{HCN}$ .

### SECTION 8.3 BONDING: THE ORBITAL

1. Which orbital has two lobes, one above and one below the nucleus? Is it a bonding orbital or an antibonding orbital?

2. Which orbital has two lobes, one to the left and one to the right of the nucleus? Is it a bonding orbital or an antibonding orbital?

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**A Loxley**



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