

The following is a summary of most of the concepts, relationships, topics that you should know for this course. Note that this may include typographical errors because it has not been proof read and is a first draft, so please check all expressions with those in the book. I would only use this as a study guide and not as a replacement for studying the topic in depth within the book. Also note that I provide information about certain things I recommend you know how to do, calculate, derive, etc. *Prof. Musgrave*

Section 1.2a – Perfect gas law

Calculating relationships between p , V and T at one state and another.

The molecules in perfect gases experience no interactions.

Do not apply the perfect gas law to systems where it does not apply (liquids, solids, dense gases)

Section 1.2b – Kinetic model of gases

Calculating average speeds of gas molecules as a function of M and T .

Section 1.2c – Mixtures of gases

Calculating partial pressures from total pressure and composition (mole fractions)

Dalton's Law: The sum of the pressures of gases of each component alone in a vessel sum to the total pressure of the mixture of those gases.

Section 1.3a – The compression factor

The volumes of real gases are different from perfect gases by the compression factor.

The compression factor is related to the Virial equation of state

Section 1.4a – van der Waals equation

Corrections for the size of molecules and their interactions can be included to form a more accurate description of the equation of state $p(V,T)$

There is no closed form expression of $V(p,T)$ – solve for V numerically.

There are closed forms of $p(T,V)$ and $T(p,V)$ so inverting partial derivatives is sometimes convenient to determine derivatives of V with respect to T or p .

Section 1.4b – van der Waals equation features

The vdW equation of state differs from the ideal gas law at high p and low T (where the assumption of ideal behavior breaks down).

Chapter 1: Key Equations:

Perfect gas law: $pV = nRT$

Temperature scales: $T(K) = T(C) + 273.15$

Partial pressures: $p = \sum p_i$ and $p_i = x_i p = \frac{n_i}{n} p$

Van der Waals Equation of state: $p = nRT / (V - nb) - an^2 / V^2$

Section 2.1 – Work, heat and energy

The energy of a system is its capacity to do work

Work is the transfer of energy by moving against a force – it is the transfer of energy by organized motion.

Heat is the transfer of random, thermal motion. It transfers disordered energy.

An adiabatic process transfers no heat and thus no random energy or entropy.

An isothermal process involves no change in T .

Section 2.2 – Internal energy U

Internal energy (U) is the total energy of the system. It includes kinetic energy and potential energy and is a state of the system. It increases because either work is done on the system or heat is transferred to the system.

Section 2.2a – The internal energy of a perfect gas depends on T only – $U(T)$.

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Progress in Physical Chemistry Volume 3 Franz Michael Dolg, 2011-11-22 Progress in Physical Chemistry is a collection of recent Review Articles published in the Zeitschrift f r Physikalische Chemie The third volume of the series Progress in Physical Chemistry comprises 27 articles most of them with review character written by the members of the Priority Program SPP 1145 of the German Research Foundation DFG

Modern Alchemy Mark Morrisson, 2007-04-19 Alchemists are generally held to be the quirky forefathers of science blending occultism with metaphysical pursuits Although many were intelligent and well intentioned thinkers the oft cited goals of alchemy paint these antiquated experiments as wizardry not scientific investigation Whether seeking to produce a miraculous panacea or struggling to transmute lead into gold the alchemists radical goals held little relevance to consequent scientific pursuits Thus the temptation is to view the transition from alchemy to modern science as one that discarded fantastic ideas about philosophers stones and magic potions in exchange for modest yet steady results It has been less noted however that the birth of atomic science actually coincided with an efflorescence of occultism and esoteric religion that attached deep significance to questions about the nature of matter and energy Mark Morrisson challenges the widespread dismissal of alchemy as a largely insignificant historical footnote to science by prying into the revival of alchemy and its influence on the emerging subatomic sciences of the late 19th and early 20th centuries Morrisson demonstrates its surprising influence on the emerging subatomic sciences of the late 19th and early 20th centuries Specifically Morrisson examines the resurfacing of occult circles during this time period and how their interest in alchemical tropes had a substantial and traceable impact upon the science of the day Modern Alchemy chronicles several encounters between occult conceptions of alchemy and the new science describing how academic chemists inspired by the alchemy revival attempted to transmute the elements to make gold Examining scientists publications correspondence talks and laboratory notebooks as well as the writings of occultists alchemical tomes and science fiction stories he argues that during the birth of modern nuclear physics the trajectories of

science and occultism so often considered antithetical briefly merged
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