A. Covalent & Ionic Compounds/Bonds

When an atom loses an electron it becomes a ____positive ion_____ (cation),

and when it gains an electron it becomes a ____negative ion_____ (anion).

 An element in the halogen group has _7__ valence electrons, and wants to ___gain 1 more to have S. This is because of the __octet____ rule.

3. An _ionic____ bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal .

In a covalent bond, atoms ____share__ electrons, whereas in an ionic bond, atoms transfer electrons.

5. (Honors only) When electrons are shared unequally in a covalent bond it is _________, and when they're shared equally it is _______nonpolar______. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the ______electronegativities______.

 An atom can have up to _8___ valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called ____ noble gases____.

7. Atoms can gain, lose, or share ___valence electrons____ to become more stable.

B. Formulas for Ionic compounds - Fill in the chart below

and the second s	Magnesium	Fluorine	Nitrogen	Lithium
# OF VALENCE ELECTRONS	2	7	5	1
OXIDATION #	+2	-1	-3	+1

2. What is the **chemical formula** for each of the following ionic compounds? (hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.)

- a) Potassium and Chlorine ____KCl___
- b) Calcium and Bromine ____CaBr2____
- c) Magnesium and oxygen ____MgO___

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SA Adler

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Table of Contents Physical Science Chemical Bonds Study Guide

- 1. Understanding the eBook Physical Science Chemical Bonds Study Guide
 - The Rise of Digital Reading Physical Science Chemical Bonds Study Guide
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Physical Science Chemical Bonds Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - $\circ\,$ Features to Look for in an Physical Science Chemical Bonds Study Guide
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Physical Science Chemical Bonds Study Guide
 - Personalized Recommendations
 - Physical Science Chemical Bonds Study Guide User Reviews and Ratings
 - Physical Science Chemical Bonds Study Guide and Bestseller Lists
- 5. Accessing Physical Science Chemical Bonds Study Guide Free and Paid eBooks
 - Physical Science Chemical Bonds Study Guide Public Domain eBooks
 - Physical Science Chemical Bonds Study Guide eBook Subscription Services
 - Physical Science Chemical Bonds Study Guide Budget-Friendly Options
- 6. Navigating Physical Science Chemical Bonds Study Guide eBook Formats

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- Physical Science Chemical Bonds Study Guide Enhanced eBook Features
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- 9. Balancing eBooks and Physical Books Physical Science Chemical Bonds Study Guide
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Physical Science Chemical Bonds Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Physical Science Chemical Bonds Study Guide
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 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Physical Science Chemical Bonds Study Guide
 - Fact-Checking eBook Content of Physical Science Chemical Bonds Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
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