

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

Na, K, S, O, Cl

3. Which atom has the higher electronegativity: sodium or sulfur?

Sulfur (sulfur is to the right of sodium and closer to fluorine)

4. Which two elements in the same family generally have similar properties?

There have been two cases here where a strong positive correlation exists

5. Arrange **fluorine** the following properties (increased or decreased) from left to right across the periodic table:

- a. atomic radius (increasing to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Seventh period, francium (bottom left corner of the periodic table)

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

Second period, fluorine (top right corner of the periodic table)

8. Write the atoms in each pair that has the largest atomic radius.

a. **Br** or **Cl**
 b. **Na** or **K**
 c. **Li** or **Cl**
 d. **Ca** or **Br**
 e. **Li** or **K**
 f. **Na** or **Br**

g. **Na** or **Br**
 h. **Ca** or **Br**
 i. **Cl** or **Br**
 j. **Br**

9. Circle the atom in each pair that has the greatest electronegativity.

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