

Answers

$Q = mc\Delta T$, where Q = heat energy, m = mass, and ΔT = change in temp.
Remember, $\Delta T = (T_{\text{final}} - T_{\text{initial}})$. Show all work and proper units.

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{1086.75}{15.75(175-25)} = 0.46 \text{ J/g}^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = mC(T_f - T_i) = 10.0\text{g} (0.90\text{J/g}^\circ\text{C})(55-22) = 297 \text{ J}$$

3. Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{67500 \text{ J}}{(1500 \text{ g})(57-32)} = 1.8 \text{ J/g}^\circ\text{C}$$

4. 100.0 g of 4.0°C water is heated until its temperature is 37°C. Calculate the amount of heat energy needed to cause this rise in temperature.

$$Q = mC(T_f - T_i) = 100\text{g}(4.184\text{J/g}^\circ\text{C})(37 - 4) = 14000 \text{ J}$$

5. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the specific heat capacity of mercury.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{455 \text{ J}}{(25\text{g})(155-25)} = 0.14 \text{ J/g}^\circ\text{C}$$



Physical Science If8767 Answer Key Heat Calculations

Wolfgang Guggemos



Physical Science If8767 Answer Key Heat Calculations:

Physical Science If8767 Answer Key Heat Calculations Book Review: Unveiling the Magic of Language

In a digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is really remarkable. This extraordinary book, aptly titled "**Physical Science If8767 Answer Key Heat Calculations**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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