

# Annotated periodic table of the elements

## Group names\*

- 1 Hydrogen & the alkali metals
- 2 Alkaline earth metals
- 3 Coinage metals
- 12 Volatile metals
- 13 Borogens
- 14 Chalcogens
- 15 Pnictogens
- 16 Chalcogens
- 17 Halogens
- 18 Helium group

\*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

# Annotated periodic table of the elements

## Group names\*

1 Hydrogen & the alkali metals	14 Crystallogens
2 Alkaline earth metals	15 Prochogens
11 Coinage metals	16 Chalcogens
12 Volatile metals	17 Halogens
13 Ionogens	18 Helium group

\*Groups 1 to 10 are named after their

first members i.e. Group 1 is the Lanthanum Group

Most common oxidation states, electronegativity (after Pauling)

Atomic number and symbol  
Atomic weight and name

1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
1 H 1.01 Hydrogen	2 He 4.003 Helium											3 B 10.81 Boron	4 C 12.01 Carbon	5 N 14.01 Nitrogen	6 O 16.00 Oxygen	7 F 18.99 Fluorine	8 Ne 20.18 Neon
9 Li 6.94 Lithium	10 Be 9.012 Beryllium											11 Al 26.98 Aluminum	12 Si 28.09 Silicon	13 P 30.97 Phosphorus	14 S 32.06 Sulfur	15 Cl 35.45 Chlorine	16 Ar 39.95 Argon
17 K 39.10 Potassium	18 Ca 40.08 Calcium	19 Sc 44.96 Scandium	20 Ti 47.87 Titanium	21 V 50.94 Vanadium	22 Cr 52.00 Chromium	23 Mn 54.94 Manganese	24 Fe 55.85 Iron	25 Co 58.93 Cobalt	26 Ni 58.69 Nickel	27 Cu 63.55 Copper	28 Zn 65.38 Zinc	29 Ga 69.72 Gallium	30 Ge 72.63 Germanium	31 As 74.92 Arsenic	32 Se 78.96 Selenium	33 Br 79.90 Bromine	34 Kr 83.80 Krypton
37 Rb 85.47 Rubidium	38 Sr 87.62 Strontium	39 Y 88.91 Yttrium	40 Zr 91.22 Zirconium	41 Nb 92.91 Niobium	42 Mo 95.94 Molybdenum	43 Tc 98.91 Technetium	44 Ru 101.07 Ruthenium	45 Rh 101.07 Rhodium	46 Pd 106.42 Palladium	47 Ag 107.87 Silver	48 Cd 112.41 Cadmium	49 In 114.82 Indium	50 Sn 118.71 Tin	51 Sb 121.76 Antimony	52 Te 127.60 Tellurium	53 I 126.91 Iodine	54 Xe 131.29 Xenon
55 Cs 132.91 Cesium	56 Ba 137.33 Barium	57 to 70 Lanthanum	71 La 138.91 Lanthanum	72 Ce 140.12 Cerium	73 Pr 140.91 Praseodymium	74 Nd 144.24 Neodymium	75 Pm 144.91 Promethium	76 Sm 150.36 Samarium	77 Eu 151.96 Europium	78 Gd 157.25 Gadolinium	79 Tb 158.93 Terbium	80 Dy 162.50 Dysprosium	81 Ho 164.93 Holmium	82 Er 167.26 Erbium	83 Tm 168.93 Thulium	84 Yb 173.05 Ytterbium	85 Lu 174.96 Lutetium
87 Fr 223.02 Francium	88 Ra 226.03 Radium	89 to 102 Actinium	103 Ac 227.03 Actinium	104 Th 232.04 Thorium	105 Pa 231.04 Protactinium	106 U 238.03 Uranium	107 Np 237.05 Neptunium	108 Pu 244.06 Plutonium	109 Am 243.06 Americium	110 Cm 247.07 Curium	111 Bk 247.07 Berkelium	112 Cf 251.08 Californium	113 Es 252.08 Einsteinium	114 Fm 257.09 Fermium	115 Md 258.10 Mendelevium	116 No 259.10 Nobelium	117 Ts 289.10 Tennessine
			103 La 138.91 Lanthanum	104 Ce 140.12 Cerium	105 Pr 140.91 Praseodymium	106 Nd 144.24 Neodymium	107 Pm 144.91 Promethium	108 Sm 150.36 Samarium	109 Eu 151.96 Europium	110 Gd 157.25 Gadolinium	111 Tb 158.93 Terbium	112 Dy 162.50 Dysprosium	113 Ho 164.93 Holmium	114 Er 167.26 Erbium	115 Tm 168.93 Thulium	116 Yb 173.05 Ytterbium	117 Lu 174.96 Lutetium
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Most common oxidation states, electronegativity (Allen-Pauling)

Atomic number and symbol, atomic weight and name

57 to 70

89 to 102

## CLASSIFICATION

Metals	Nonmetals
Pre-transition	
Active	Post-transition
Alkaline earth	
Alkaloids	
Transition	Intermediate
Post-transition	
Metals	
Metals (Ru-Pt, Os-Au)	Metals gas

## Active metals

### Pre-transition metals

Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Fe, Mg and Al have any structural uses. Most form basic oxides (Be and B are amphoteric).

### Rare earths

Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-colored compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.

### Actinides

Soft, dense and reactive metals, those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form colored compounds.

## Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

### Noble metals

Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalysts properties. Very weakly electropositive; inert to acids.

### Frontier metals

Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterized, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

## Metals

## Nonmetals

### Metalloids

Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

### Intermediate nonmetals

B and Si are colorless gases, C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies; low to high electron affinities; moderate to high electronegativity values; relatively poor to moderately strong oxidising agents; tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (P<sub>2</sub>O, CO, NO, N<sub>2</sub>O).

\*as black P, the most stable form in ambient conditions

### Group 1

The starts as Li-Li, and then bifurcates into a La branch i.e. La-Ac, and a Lu branch i.e. Lu-Lu. Trends going down: Li-Li, and then La branch, resemble those of groups 1–2.

• Lu branch, more resemble those of groups 1–2

## Radioactive

Approximation is likely to be either a metalloid or a post-transition metal

## Corrosive nonmetals

O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid, I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.

## Noble gases

Colorless, colorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePF<sub>6</sub>, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

# Periodic Table Unscrambled Table Answers

**Anthony S. Fauci**



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