Annotated periodic table of the elements

Group names*

1 Hydrogen & the alkali metals
2 Alkaline earth metals
11 Coinage metals
12 Volatile metals
13 Icosagens
14 Crystallogens
15 Pnictogens
16 Chalcogens
17 Halogens
18 Helium group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium Group

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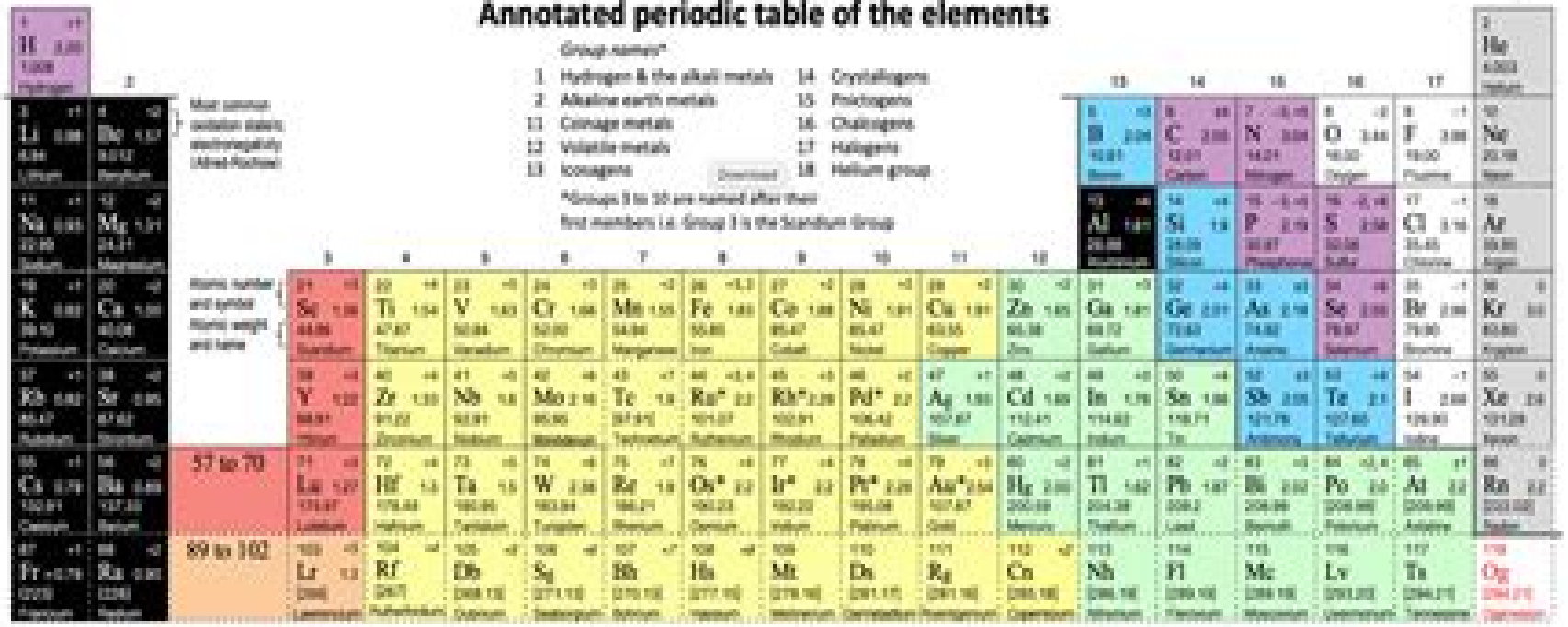
Most common oxidation states; electronegativity (Allred-Rochow)

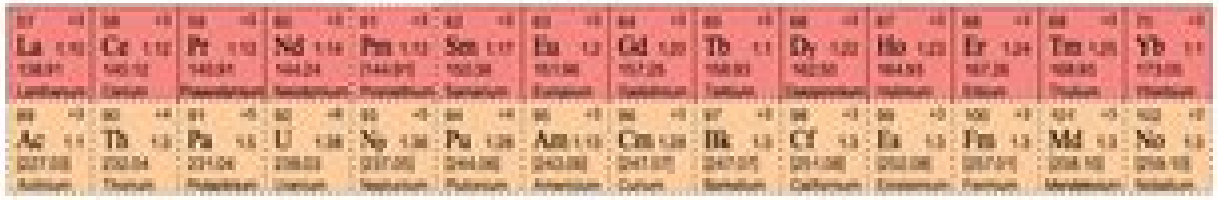
Atomic number and symbol; Atomic weight and name

57 to 70

89 to 102



CLASS COUNTERPARTS		
Metals		Nonmetals
Active	Pre-transition	Corrosive
	Rare earth	
	Actinide	
Transition		Intermediate
Frontier		Metalloid
Noble (Ru–Pt; Os–Au)		Noble gas



Radioactive

Oganesson is likely to be either a metalloid or a post-transition metal

Metals | Nonmetals

Active metals

Pre-transition metals
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).

Rare earths
Closely resemble Ca, Sr and Ba, but many are heavier, and they form mostly pale-coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Y& and Lu exhibit weak amphoterism.

Actinides
Soft, dense and reactive metals, those from Es onwards are synthetic. Early actinides (Th–Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Rf onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.

Noble metals
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very weakly electropositive; insoluble oxides.

Frontier metals
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or fewer nearest neighbours than other metals. All used or trialed as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphoterism, and anionic species such as zincates, stannates, and bismuthates.

Metalloids

Metallic looking, brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.9–2.2). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.

Intermediate nonmetals

H and N are colorless gases; C, P*, and Se are metallic looking solids while S is yellow. The latter four are brittle but all known in malleable, pliable or ductile forms. Moderate to high ionisation energies, low to high electron affinities; moderate to high electronegativity values, relatively poor to moderately strong oxidising agents; tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H_2O, CO, NO, N_2O).

* as black P, the most stable form in ambient conditions.

Group 1
This starts as Sc-Y-, and then bifurcates into a Lu branch i.e. -Lu-Ac, and a La branch i.e. -La-Lr. Trends going down Sc-Y- and the:

- Lu branch, resemble those of groups 1-2
- La branch, more resemble those of groups 4-10

Corrosive nonmetals

O, F, and Cl are, respectively, colorless, pale yellow, and pale yellow-green gases; Br is a red-brown liquid; I is a metallic looking solid. Chemically tend to have to high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominately ionic compounds with metals. Their oxides are acidic.

Noble gases
Colorless, odorless, non-flammable gases with very low chemically reactivity. The first noble gas compound, of approximate composition $XePtF_6$, was not prepared until 1962. More than half a thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

February 4th, 2020

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Victor M. Corman

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