

# 6

## Chemical Reactions

We have discussed a large number of chemical reactions. Some are accompanied by a change in colour, some by the evolution of a gas, some by the formation of a precipitate, and so on. At the molecular level, different types of changes occur. In some reactions, smaller parts add up to form bigger entities, whereas in some others, bigger entities break into smaller ones. Again, in some reactions, one element displaces another from a compound whereas in some others, compounds exchange radicals among themselves. For a systematic study, therefore, it is essential that we classify the reactions into different types.

However, we will discuss first the change in energy in chemical reactions, which is common to all types.

### Change in Energy in a Reaction

All reactions are accompanied by a change in energy. Energy (in the form of heat or light) is either given out or taken in during a chemical reaction.

A chemical reaction in which heat or light is evolved is called an exothermic reaction.

A chemical reaction in which heat or light is absorbed is called an endothermic reaction.

In an exothermic process, heat or light is given out to the surroundings. That is why you feel hot when you accidentally touch anything burning. On the other hand, in an endothermic process, heat or light is taken

in, i.e., absorbed from the surroundings. This is why a cube of ice is cold to touch. The ice tends to melt when you touch it. The melting of ice is an endothermic process, and the heat required for it is drawn from your hand, and so the cold feeling. Remember that the melting of ice is a physical change, not a chemical reaction.

### Why a Change in Energy in a Reaction

As we know, any two atoms in a molecule are held together by a force of attraction called a chemical bond. And also that a chemical bond is much stronger than an intermolecular force. We also know that it is the atoms that take part in a chemical reaction.

Thus, we can consider a chemical reaction to be a result of the following phenomena:

- The breaking of bonds of the reactant molecules so as to set the atoms free for a new combination. (Energy is absorbed in the process.)
- The forming of fresh bonds between the new partners so that the product molecules are formed. (Energy is given out in the process.)

The energy required for the breaking of the bonds of the reactant molecules is not the same as that given out in the formation of the bonds between the new partners. When the energy given out in the formation of the bonds is more than that used for the breaking of the bonds, the surplus energy is evolved and the process is exothermic.

However, the breaking of bonds is given out in difference in for from out be absorbed reaction to endothermic.

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web jul 15 2021 certains des noms de drogues les plus courants pour les stimulants sont les suivants caféine nicotine amphétamines cocaïne 2 déprimeurs les effets de ces drogues déprimeurs incluent la réduction de la sensation de tension ainsi que le soulagement de l anxiété et la relaxation musculaire

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web les jeunes restent le groupe le plus susceptible de faire usage de drogues en 2021 au niveau mondial la prévalence annuelle de l usage de cannabis chez les jeunes de 15 et 16 ans était de 5 34 contre 4 3 chez les adultes en outre la consommation de drogues est particulièrement nocive pour les jeunes

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web une drogue est un composé chimique biochimique ou naturel capable d altérer une ou plusieurs activités neuronales et ou de perturber les communications neuronales la consommation de drogues par l homme afin de modifier ses fonctions physiologiques ou psychiques ses réactions physiologiques et ses états de conscience n est pas récente

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web on appelle drogue toute substance psychotrope ou psychoactive qui perturbe le fonctionnement du système nerveux central sensations perceptions humeurs sentiments motricité ou qui modifie les états de conscience une drogue est un produit susceptible d entraîner une dépendance physique et ou psychique

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