

15 Water and Aqueous Systems



BONDING AND INTERACTIONS

15.1 Water and Its Properties



For students using the Foundation edition, assign problems 1–7, 9, 11, 12.

Essential Understanding

The bonding between water molecules differs when water is in different states, giving it different properties.



Reading Strategy

Compare and Contrast Organizing information in a table helps you compare and contrast several topics at one time. Compare water in the liquid state to water in the solid state. As you read, ask yourself, “How are they similar? How are they different?”

As you read Lesson 15.1, use the compare and contrast table below.

Attributes of water	Water in the liquid state	Water in the solid state
Physical properties	<i>liquid; flows; colorless</i>	<i>hard; does not flow; colorless</i>
How atoms bond	<i>hydrogen bonds are irregular</i>	<i>hydrogen bonds form hexagons</i>
Relative density	<i>more dense</i>	<i>less dense</i>

EXTENSION Make a Venn diagram of the information in the compare and contrast table.



Lesson Summary

Water in the Liquid State The polarity of water molecules causes its high surface temperature, low vapor pressure, and high boiling point.

- ▶ The O—H bonds in a water molecule are polar because oxygen is more electronegative than hydrogen.
- ▶ Hydrogen bonding between water molecules occurs because a hydrogen atom is weakly bonded to an unshared electron pair of another water molecule's oxygen atom.

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KJ Lindholm-Leary



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