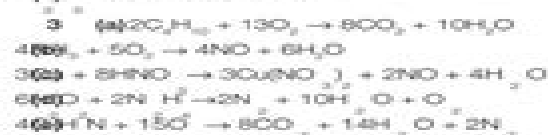
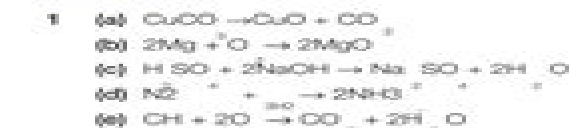


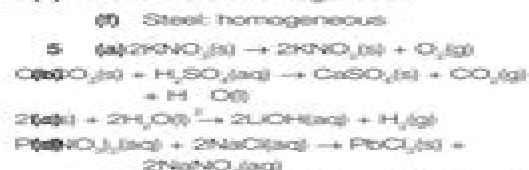
Answers

Chapter 1

Exercises



4 (a) Sand and water: heterogeneous
 (b) Sugar and water: homogeneous
 (c) Oil and iron filings: heterogeneous
 (d) Ethanol and water: homogeneous



In the answers, exercises and text there is inconsistency about whether the water formed in combustion reactions is gas or liquid. At room temperature it will ultimately finish up as a liquid, as shown here.

6 X has diffused more quickly, so it must be a lighter gas. Its particles have greater velocity than the particles of Y at the same temperature. (Note though that they will both have the same value for average kinetic energy.)

7 From the kinetic molecular theory we would expect a solid to be more dense than its liquid, and therefore that ice would sink in water.

8 Bubbles will be present through the volume of the liquid. A brown gas is visible above the brown liquid. As the two states are at the same temperature, the particles have the same average kinetic energy and are moving at the same speed. The inter-particle distances in the gas are significantly larger than those in the liquid.

9 At certain conditions of low temperature and low humidity, snow changes directly to water vapour by sublimation, without going through the liquid phase.

10 Steam will condense on the skin, releasing energy as it forms liquid at the same temperature (E–D on Figure 1.4). This is additional to the energy released when both the boiling water and the condensed steam cool on the surface of the skin.

11 B

12 AW 01_EX_12

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1. Understanding the eBook Pearson Chemistry Worksheet Answers
 - The Rise of Digital Reading Pearson Chemistry Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Chemistry Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pearson Chemistry Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Chemistry Worksheet Answers
 - Personalized Recommendations
 - Pearson Chemistry Worksheet Answers User Reviews and Ratings
 - Pearson Chemistry Worksheet Answers and Bestseller Lists
5. Accessing Pearson Chemistry Worksheet Answers Free and Paid eBooks
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- Pearson Chemistry Worksheet Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Pearson Chemistry Worksheet Answers
 - Interactive Elements Pearson Chemistry Worksheet Answers
- 8. Staying Engaged with Pearson Chemistry Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pearson Chemistry Worksheet Answers
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pearson Chemistry Worksheet Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Setting Reading Goals Pearson Chemistry Worksheet Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Chemistry Worksheet Answers
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