

16.4**CALCULATIONS INVOLVING
COLLIGATIVE PROPERTIES****Section Review****Objectives**

- Calculate the molality and mole fraction of a solution.
- Describe how the freezing-point depression and boiling-point elevation are related to molality.

Vocabulary

- molality (m)
- mole fraction
- molal freezing-point depression constant (K_f)
- molal boiling-point elevation constant (K_b)

Key Equations

- Molality = $\frac{\text{moles of solute}}{\text{kilogram of solvent}}$
- mole fractions: $X_A = \frac{n_A}{n_A + n_B}$ $X_B = \frac{n_B}{n_A + n_B}$
 where n_A = moles of solute
 n_B = moles of solvent
- $\Delta T_b = K_b \times m$
- $\Delta T_f = K_f \times m$

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

Molality is an expression of concentration involving the ratio of _____ 1 _____ particles to _____ 2 _____ particles. Molality is expressed as moles of solute per _____ 3 _____ of solvent.

Another expression of concentration is _____ 4 _____, in which concentrations are expressed as the ratio of moles of solute to the total number of moles of solvent and solute. Each solvent has a characteristic _____ 5 _____ elevation constant and molal freezing-point _____ 6 _____ constant. The elevation in boiling point of a solution can be calculated by multiplying the _____ 7 _____ concentration of the solution by the boiling-point _____ 8 _____ constant of the solvent.

- _____
- _____
- _____
- _____
- _____
- _____
- _____
- _____

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